

CHAPTER 5 REVIEW

The Periodic Law**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ In the modern periodic table, elements are ordered
 - (a) according to decreasing atomic mass.
 - (b) according to Mendeleev's original design.
 - (c) according to increasing atomic number.
 - (d) based on when they were discovered.
2. _____ Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing
 - (a) density.
 - (b) reactivity.
 - (c) atomic number.
 - (d) atomic mass.
3. _____ The modern periodic law states that
 - (a) no two electrons with the same spin can be found in the same place in an atom.
 - (b) the physical and chemical properties of an element are functions of its atomic number.
 - (c) electrons exhibit properties of both particles and waves.
 - (d) the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.
4. _____ The discovery of the noble gases changed Mendeleev's periodic table by adding a new
 - (a) period.
 - (b) series.
 - (c) group.
 - (d) level.
5. _____ The most distinctive property of the noble gases is that they are
 - (a) metallic.
 - (b) radioactive.
 - (c) metalloid.
 - (d) largely unreactive.
6. _____ Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4.
 - (b) 10.
 - (c) 11.
 - (d) 18.
7. An isotope of fluorine has a mass number of 19 and an atomic number of 9.
 - _____ a. How many protons are in this atom?
 - _____ b. How many neutrons are in this atom?
 - _____ c. What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

Periodic Law Answer Sheet

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 - The Rise of Digital Reading Periodic Law Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Law Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Answer Sheet
 - Personalized Recommendations

- Periodic Law Answer Sheet User Reviews and Ratings
- Periodic Law Answer Sheet and Bestseller Lists
- 5. Accessing Periodic Law Answer Sheet Free and Paid eBooks
 - Periodic Law Answer Sheet Public Domain eBooks
 - Periodic Law Answer Sheet eBook Subscription Services
 - Periodic Law Answer Sheet Budget-Friendly Options
- 6. Navigating Periodic Law Answer Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Law Answer Sheet Compatibility with Devices
 - Periodic Law Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Law Answer Sheet
 - Highlighting and Note-Taking Periodic Law Answer Sheet
 - Interactive Elements Periodic Law Answer Sheet
- 8. Staying Engaged with Periodic Law Answer Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Law Answer Sheet
- 9. Balancing eBooks and Physical Books Periodic Law Answer Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Law Answer Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Law Answer Sheet
 - Setting Reading Goals Periodic Law Answer Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Law Answer Sheet
 - Fact-Checking eBook Content of Periodic Law Answer Sheet

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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