

# Percentage Yield Lab

**Purpose:** To determine the theoretical, actual and % yield of copper produced by the following reaction:



(solid) (solution)  
(solid)

**Materials and Method:** Refer to Percentage Yield lab

**Observations:**

Mass of clean, empty beaker  $\rightarrow$  78.5g

Mass of iron with beaker  $\rightarrow$  79.255g

Filter paper  $\rightarrow$  0.971g

Mass of copper inside filter paper  $\rightarrow$  3.00g

**Conclusions:**

- 1) The mass of iron used is 0.755g because 79.255g (the weight of the iron filled beaker) - 78.5g (the weight of the empty beaker) = 0.755g
- 2) Fe + Cu
- 3) Two possible reasons which could explain the difference between your actual yield and the theoretical yield are the scale not being accurate and not being read. Have we waited long enough?

# Percentage Yield Lab Grade 12

**G Orfield**



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