

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant

Excess: All reactants except the limiting reagent will have some left over. Excess is the amount of reactant that is not completely consumed.

Remember: Limiting reagent will determine how much product is produced and

byproducts will be. The limiting reagent is also a reactant.



The ratio required for the composition for the limiting reagent.

2H ₂ (g)	100.00 percent
O ₂ (g)	50.00 percent
2H ₂ O(l)	100.00 percent
H ₂ O	100.00 percent

1. 100.00 g of each element will be used & react to form 100.00 g of product water. 100.00 g of hydrogen & 50.00 g of oxygen will be used. 100.00 g of water is produced.



2. 100.00 g H₂(g) is the limiting reagent. 24.75 g O₂(g) is the excess reagent.

3. 100.00 g of hydrogen will be completely consumed & 24.75 g of oxygen will be present.



4. 100.00 g of hydrogen is present in a 100.00 g of hydrogen & 24.75 g of oxygen. 100.00 g of hydrogen is present. 100.00 g of hydrogen is present in the reaction. 100.00 g of hydrogen is present.



Percent Yield And Limiting Reactants Quiz

RJ Alexander



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