

Percent, Actual, and Theoretical Yield

Suppose a chemist reacts 10.0 g of CaCO_3 with 10.0 g of HCl to produce CaCl_2 , H_2O , and CO_2 . The chemist finds that 10.0 g of CaCl_2 is produced. What is the theoretical yield of CaCl_2 in grams?

$$\begin{aligned} \text{CaCO}_3 + 2\text{HCl} &\rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2 \\ 100.09 \text{ g/mol} & \quad 72.02 \text{ g/mol} \quad 110.98 \text{ g/mol} \quad 18.02 \text{ g/mol} \quad 44.01 \text{ g/mol} \\ 10.0 \text{ g CaCO}_3 & \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g}} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{110.98 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 11.1 \text{ g CaCl}_2 \end{aligned}$$

11.1 g CaCl_2 is the theoretical yield of CaCl_2 in grams.

$$\frac{10.0 \text{ g CaCl}_2}{11.1 \text{ g CaCl}_2} \times 100 = 90.1\%$$

Suppose a chemist reacts 10.0 g of CaCO_3 with 10.0 g of HCl . The chemist finds that 10.0 g of CaCl_2 is produced. What is the theoretical yield of CaCl_2 in grams?

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A Loxley



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