

c. Standard Enthalpy of Vaporization ΔH_{vap}°

→ The enthalpy change for the vaporization of 1 mole of liquid to a gas at the boiling point and standard pressure.

Example:

$$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g) \quad \Delta H_{vap}^\circ = +44.0 \text{ kJ/mol}$$

d. Standard Enthalpy of Melting ΔH_{fus}°

→ The enthalpy change for melting 1 mol of a solid into a liquid at the melting point and standard pressure.

Example:

$$\text{H}_2\text{O}(s) \rightarrow \text{H}_2\text{O}(l) \quad \Delta H_{fus}^\circ = +6.01 \text{ kJ/mol}$$

e. Standard Neutralization Enthalpy Change ΔH_{neut}°

→ The enthalpy change in the neutralization of an acid (H⁺) by a base (OH⁻) forming 1 mole of water.

Example:

$$\text{HCl}(aq) + \text{NaOH}(aq) \rightarrow \text{NaCl}(aq) + \text{H}_2\text{O}(l) \quad \Delta H_{neut}^\circ = -57.3 \text{ kJ/mol}$$

f. Standard Enthalpy of Formation ΔH_{form}°

→ The enthalpy change for forming 1 mol of elemental atoms in the gas phase at standard conditions.

Example:

$$\text{H}_2 \rightarrow 2\text{H} \quad \Delta H_{form}^\circ = +460 \text{ kJ/mol}$$

g. Standard Solution Enthalpy Change ΔH_{sol}°

→ The enthalpy change when 1 mole of a compound is converted into a solution at standard conditions.

Example:

$$\text{NaOH}(s) \rightarrow \text{NaOH}(aq) \quad \Delta H_{sol}^\circ = -42.6 \text{ kJ/mol}$$

Note Taking Guide For Thermochemical Equations

Christian G. Meyer



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