

Chemical Formula

Telling what provides molecular and relative of each element (in a compound)

Molecular $\text{C}_{10}\text{H}_{18}\text{O}$ (molecular)

Formula represents the number of atoms of each
in a single molecule

Empirical $\text{C}_5\text{H}_9\text{O}$ (molecular)

Formula represents the simplest ratio of
positive and negative ions in a formula
unit

Atomic Mass

The average mass of atoms measured relative
to hydrogen

Isotopes

$^{12}_6\text{C}$ stable isotope carbon

$^{13}_6\text{C}$ stable isotope carbon

$^{14}_6\text{C}$ radioactive

Abundance

$^{12}_6\text{C}$ abundant

$^{13}_6\text{C}$ stable isotope carbon

$^{14}_6\text{C}$ radioactive

Note Taking Guide Episode 601 Chemical Formula Key

Wolfgang Guggemos



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