

PAPER-I  
COMPUTER OPERATOR & PROGRAMMING ASSISTANT  
(THEORY & EMPLOYABILITY SKILLS)  
SEMESTER - I

2018  
CapA + ES

TOTAL TIME : 3 Hrs.

TOTAL MARKS: 80

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Note:- This paper contains two parts – Part A & Part B.  
Attempt all the questions.  
All questions carry equal marks.

PART-A (Theory)		(30)
1	Marks :1	1 Marks :1
CPU stands for.....		CPU का अर्थ है .....
A Common Processing Unit		A Common Processing Unit
B Central Process Unit		B Central Process Unit
C Central Processing Unit		C Central Processing Unit
D Central Performance Unit		D Central Performance Unit
2	Marks :1	2 Marks :1
The hexadecimal number consist of .....		हेक्साडिमल नंबर में ..... संख्या होती है
A 0 - 7		A 0 - 7
B 0 - 9		B 0 - 9
C 0 - 9, A - F		C 0 - 9, A - F
D 1 - 9, A - F		D 1 - 9, A - F
3	Marks :1	3 Marks :1
The memory device in the third generation computers is.....		तीसरी पीढ़ी कंप्यूटर की मेमोरी डिवाइस ... है
A Vacuum tubes		A Vacuum tubes
B Transistors		B Transistors
C Integrated circuit		C Integrated circuit
D VLSI		D VLSI
4	Marks :1	4 Marks :1
The Process of isolating and correcting the errors is called.....		त्रुटियों को खोजने और सुधारने की प्रक्रिया को ..... कहा जाता है
A program testing		A program testing
B program debugging		B program debugging

# Sample Question Paper Of Iti Entrance 2013

**D Siedentop**



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web equilibrium questions a covalent bond has a standard gibbs free energy of  $400 \text{ kJ mol}^{-1}$  while the hydrogen bond and van der waals forces have substantially lower values respectively  $25$  and  $5 \text{ kJ mol}^{-1}$  given such  $\Delta G$  values which of the following is equivalent to the ratio of the  $K_{eq}$  value of a covalent bond to that of a hydrogen bond

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the likely result of a chemical reaction in the previous chapter we discussed the principles of chemical kinetics which deal with the rate of change or how quickly a given chemical reaction occurs

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web nov 13 2022 define the equilibrium state of a chemical reaction system what is its practical significance state the meaning and significance of the following terms reversible reaction quantitative reaction kinetically inhibited reaction explain the meaning of the statement equilibrium is macroscopically static but microscopically dynamic very

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web in a chemical equilibrium the forward and reverse reactions occur at equal rates and the concentrations of products and reactants remain constant if we run a reaction in a closed system so that the products cannot escape we often find the reaction does not give a 100 yield of products