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## **MECHANICAL ENGINEERING GROUP**

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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web limiting reagents by caroline gumer home using limiting reagents percentage yield answer key percentage yield 1 for the balanced equation shown below if the reaction of 90.6 grams of  $\text{CO}$  produces 36.7 grams of  $\text{C}_3\text{H}_8$  what is the percent yield  $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}$  what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3 \text{ g}}{32 \text{ g/mol}} = 0.666 \text{ mol}$

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 160.0 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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