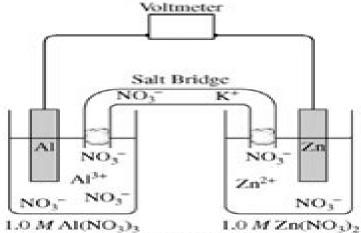
AP Chemistry Chapter 20 Electrochemistry Practice Free Response #1

$$2 \text{ Al}(z) + 3 \text{ Zn}^{2*}(ag) \rightarrow 2 \text{ Al}^{3*}(ag) + 3 \text{ Zn}(z)$$

Respond to the following statements and questions that relate to the species and the reaction represented above.

- (a) Write the complete electron configuration (e.g., 1x2 2x2 ...) for Zn2+
- (b) Which species, Zn or Zn2+, has the greater ionization energy? Justify your answer.
- (c) Identify the species that is oxidized in the reaction. The diagram shows a galvanic cell based on the reaction. Assume that the temperature is 25%



- (d) The diagram includes a salt bridge that is filled with a saturated solution of KNO<sub>2</sub>. Describe what happens in the salt bridge as the cell operates.
- (e) Determine the value of the standard voltage. Eo. for the cell.

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