

Practice Final Exam Chem 112

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. Three hundred milliliters (150 mL) of 4.50 M sugar solution is diluted to 600 mL by the addition of pure water. Which of these statements about the solution is correct?
a. The moles of water remain the same. b. The moles of sugar decrease. c. The concentration of sugar increases to 5.0 M. d. The moles of sugar remain constant. e. All of these statements are incorrect.
- _____ 2. How many moles of solute are present in 75.0 grams of solution that is 1.30% by mass glucose ($C_6H_{12}O_6$)? $MM(C_6H_{12}O_6) = 180.16 \text{ g/mol}$
a. 5.4×10^{-3} moles b. 0.800 moles c. 176 moles d. 1.76×10^{-3} moles e. 225 moles
- _____ 3. A concentrated reagent is sold as a solution that is 60.0 mass % solute. If the molar mass of the solute is 244 g/mol and the density of this solution is 1.20 g/mL, what is the molarity of the solution?
a. 10.0 M b. 2.95 M c. 15.5 M d. 0.72 M e. 1.45 M
- _____ 4. The boiling point of pure benzene (C_6H_6) is 80.10 °C. A solution of 14.2 g of a solid dissolved in 64.0 g of benzene boils at 81.09 °C. If k_b for benzene is 2.53 °C/molal, the molar mass of the solid is:
a. 662 g/mol b. 56.0 g/mol c. 92.0 g/mol d. 628 g/mol e. 568 g/mol
- _____ 5. What is the freezing point for a 0.71 m solution of glycerol ($C_3H_8O_3$) in ethanol? Freezing point(ethanol) = -114.6 °C; $K_f(\text{ethanol}) = 1.99 \text{ °C/m}$
a. 0.10 °C b. -113.2 °C c. -1.41 °C d. -116.0 °C e. none of these
- _____ 6. At 450 °C, $K_c = 0.231$ for the reaction:



What is the value of the equilibrium constant for the following reaction at 450 °C?



- a. 0.481 b. 0.156 c. 0.231 d. 0.053 e. 4.329

- _____ 7. Consider the following system at equilibrium and its corresponding concentration equilibrium constant:



What is the value of K_p for this reaction at 470 °C?

- a. $K_p = 3.80 \times 10^{-6}$ b. $K_p = 8.53 \times 10^{-6}$ c. $K_p = 4.32 \times 10^{-5}$ d. $K_p = 1.10 \times 10^{-4}$ e. $K_p = 1.19 \times 10^4$

- _____ 8. For the reaction $\text{I}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{ICl}(\text{g})$, the relation between K_p and K_c is:
a. $K_p = K_c$ b. $K_p = K_c(RT)^2$ c. $K_p = K_c(RT)^{-2}$ d. $K_c = K_p(RT)^2$ e. $K_p = K_c(RT)^4$

Practice General Chemistry 112 Final Exam

Yi-Tong Ma



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United States Air Force Academy United States Air Force Academy,1983 **Annual Catalogue** United States Air Force Academy,1984 **Undergraduate Announcement** University of Michigan--Dearborn,1991 **The Journal of Education** ,1927 **Bulletin of the Engineering Extension Division of the Virginia Polytechnic Institute** Virginia Polytechnic Institute. Engineering Extension Division,1924 *Engineering Extension Division Series Bulletin* Virginia Polytechnic Institute, Blacksburg. Engineering Extension Division,1929 **Engineering Extension Division Series Bulletin** Virginia Polytechnic Institute. Engineering Extension Division,1929 **Getting into Medical School For Dummies** Carleen Eaton,2013-05-03 Your plain English guide to getting into the medical school of your dreams Getting accepted to medical school is a long and rigorous process and many students find they need help If you re one of these students Getting into Medical School For Dummies is the perfect tool to help you through the process and realize your dream By providing you with concise information about preparing for and applying to medical school Getting into Medical School For Dummies prepares you for the application process Written by an industry expert it gives you a distinct advantage in the competitive medical school admissions process preparing you for every step and helping you create your best application Takes you through the often overwhelming process of applying to medical school Explains what medical schools and admissions committees are really looking for Provides plain English explanations of complicated medical school admissions processes If you re one of the over 40 000 students who apply to medical school each year and need help sorting through the admissions schedule writing statements of intent and preparing to take the MCAT Getting Into Medical School For Dummies has you covered **Library Journal** Melvil Dewey,Richard Rogers Bowker,L. Pylodet,Charles Ammi Cutter,Bertine Emma Weston,Karl Brown,Helen E. Wessells,1965 Includes beginning Sept 15 1954 and on the 15th of each month Sept May a special section School library journal ISSN 0000 0035 called Junior libraries 1954 May 1961 Also issued separately Super 10 Mock Tests for UPSC EPFO (Enforcement Officers/Accounts Officers) Exam (2020-21) Disha Experts,2020-06-20

Commonwealth Universities Yearbook ,1927 *Industrial & Engineering Chemistry* ,1928 *Booklist* ,1966 **Study Guide to Accompany Basics for Chemistry** Martha Mackin,2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that

students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

ASVAB For Dummies Rod Powers, Jennifer Lawler, 2007-05-22 Packed with practice questions and proven study tips Get fully briefed on the changes to the ASVAB and sharpen your test taking skills Want to ace the ASVAB This essential guide provides a comprehensive review of all test subjects and covers the latest updates including the new short length ASVAB and a new sample of the Armed Forces Qualifying Test You ll discover the pros and cons of the paper and computer exams which tests are important to your military career and cutting edge study techniques Understand the test s formats Prepare to take the ASVAB Improve your study techniques Memorize key concepts Conquer the subtests Compute your scores Match scores to military jobs Maximize your career choices

Corrosion in the Petrochemical Industry, Second Edition ,2015-12-01 Originally published in 1994 this second edition of Corrosion in the Petrochemical Industry collects peer reviewed articles written by experts in the field of corrosion that were specifically chosen for this book because of their relevance to the petrochemical industry This edition expands coverage of the different forms of corrosion including the effects of metallurgical variables on the corrosion of several alloys It discusses protection methods including discussion of corrosion inhibitors and corrosion resistance of aluminum magnesium stainless steels and nickels It also includes a section devoted specifically to petroleum and petrochemical industry related issues

Cumulated Index Medicus ,1989

Undergraduate and Graduate Studies South Carolina College, University of South Carolina, 1969

Million Dollar Data: Building Confidence - Vol.1 Stephen DeMeo, 2020-07-31 Global warming our current and greatest challenge is without precedent Among the many consequences that are impacting our society one unanticipated concern involves scientific truth When the President of the United States and others in his administration declare that global warming is fake science it calls into question what real science is and what real school science should be I will argue that real science is quality science one that is based on the rigorous collection of reliable and valid data To collect quality data requires bending over backwards to get things right and this is exactly what makes science so special Truth is made when scientists go this extra yard and devise controlled experiments collect large data sets confirm the data and rationally analyze their results Making scientific truth sounds difficult to do in the science laboratory but in reality there are many straightforward ways that truth can be constructed In the first of two volumes I discuss twelve such ways I call them Confidence Indicators that can allow students

to strongly believe in their data and their subsequent results Many of these methods are intuitive and can be used by young students on the late elementary level all the way up to those taking introductory college science courses As in life science is not without doubt In the second volume I introduce the concept of scientific uncertainty and the indicators used to calculate its magnitude I will show that science is about connecting confidence with uncertainty in a specific manner what I refer to as the Confidence Uncertainty Continuum expression This important relationship epitomizes the scientific enterprise as a search for probabilistic rather than absolute truth This two volume set will contain a variety of ways that data quality can be instituted into a science curriculum To support its use many of the examples that I will present involve science teachers as well as student work and feedback from different grade levels and in different scientific disciplines Specific chapters will be devoted to reviewing the academic literature on data quality as well as describing my own personal research on this important but often neglected topic

Vocational Education of Junior College Grade Guy Vernon Bennett, 1928

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