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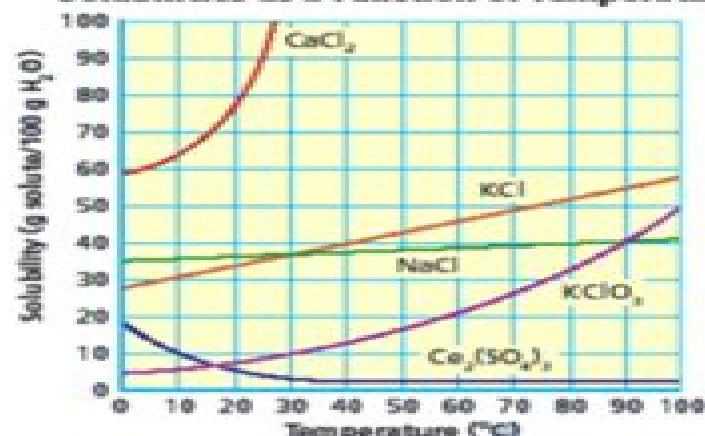
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Solubility Curve Worksheet 2



Directions: Use the graph below to answer the following questions.

Solubilities as a Function of Temperature



Types of Solutions

- Points that collectively make up the solubility curve (points **ON** the curve) represent _____ solutions.
- Points that are **BELOW** the curve represent _____ solutions.
- Points **ABOVE** the solubility curve represent _____ solutions, and the **difference** between the point above the curve and on the curve represents the amount of solute which will precipitate out.
- Say if the resulting solutions would be saturated, supersaturated or unsaturated.
 - 60 g of KCl at 70°C _____
 - 10 g of KClO₃ at 60°C _____
 - 70 g of CaCl₂ at 20°C _____

1. a. What is the solubility of calcium chloride (CaCl₂) at 5°C? _____
 b. What is the solubility of calcium chloride at 25°C? _____
 c. What is happening to the solubility and temperature for CaCl₂? Solubility is _____ as temperature is _____.
2. a. What is the solubility of cerium sulfate - Ce₂(SO₄)₃ at 10°C? _____
 b. What is the solubility of cerium sulfate at 50°C? _____
 c. What is happening to the solubility and temperature for Ce₂(SO₄)₃? Solubility is _____ as temperature is _____.
3. Which substance on the graph is **least** soluble at 10°C? _____
4. Which substance on the graph shows the **least** change in solubility from 0°C to 100°C? _____
5. Which solution is more concentrated:
 - a) At 10°C, a saturated solution of NaCl or a saturated solution of CaCl₂ (circle one)
 - b) At 55°C, a saturated solution of KClO₃ or a saturated solution of Ce₂(SO₄)₃ (circle one)
6. What is the mass of cerium sulfate that will dissolve in 50 g of water at 10 degrees Celsius?
7. What is the mass of potassium chloride that will dissolve in 250 g of water at 70 degrees Celsius?
8. At 90 degrees Celsius, 10 g of potassium chlorate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? How do you know?
9. A saturated solution of potassium chlorate is dissolved in 100g of water. If the saturated solution is cooled from 90 degrees Celsius to 60 degrees Celsius, how many grams would crystallize out?

Practice Solubility Problems

Abby S. Bloch



Practice Solubility Problems:

This book delves into Practice Solubility Problems. Practice Solubility Problems is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Practice Solubility Problems, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Practice Solubility Problems
- Chapter 2: Essential Elements of Practice Solubility Problems
- Chapter 3: Practice Solubility Problems in Everyday Life
- Chapter 4: Practice Solubility Problems in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Practice Solubility Problems. This chapter will explore what Practice Solubility Problems is, why Practice Solubility Problems is vital, and how to effectively learn about Practice Solubility Problems.

3. In chapter 2, the author will delve into the foundational concepts of Practice Solubility Problems. This chapter will elucidate the essential principles that need to be understood to grasp Practice Solubility Problems in its entirety.

4. In chapter 3, this book will examine the practical applications of Practice Solubility Problems in daily life. The third chapter will showcase real-world examples of how Practice Solubility Problems can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Practice Solubility Problems in specific contexts. This chapter will explore how Practice Solubility Problems is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Practice Solubility Problems. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Practice Solubility Problems.

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