Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use q = (m)(Cp)(AT) to solve the following problems. Show all work and units.

How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$q = ?$$
 $\Delta t = 55 - 22$ $q = m \cdot \Delta t$
 $m = 10$ $q = \frac{33}{2}$ $q = m \cdot \Delta t$ $q = 10 \times 0$

The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did 1. this sample absorb? C for water = 4.18 J/g°C

25 How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$bt = 25 - 10 = 15$$
 $m = 2259$
 $S = 4.18 5/5c$

$$a_1 = m \le 8t$$

 $= 22.5 \times 4.17 \times 15$
 $= 1410755$
 $IJ = 1 \times 10^3 \text{ KS}$ $= 14.1075 \text{ KS}$

To bring 1.0kg of water from 25°C to 99°C takes how much heat input? m= 1 14= 10000 91-1000 X 4-18 X 74

An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal?

A calorimeter has a heat capacity of 1265 J/C. A reaction causes the temperature of the 5. calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process? C = 1265 5/c.

$$C = 1265 \text{ J/c}$$

$$\Delta t = 25.12 - 22.34 = 2.78$$

$$Q = ?$$

$$Q = C\Delta t$$

$$= 1265 \times 2.78$$

$$= 3516.75$$

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Silberberg

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