#### EMPIRICAL AND MOLECULAR FORMULA

- What is the empirical formula of a compound that contains 46.3 % lithium and 53.7% oxygen?
- 2) What is the empirical formula of a compound that contains 15.9 % boron and 84.1 % fluorine?
- 3) Phosphorus reacts with oxygen to give a compound that is 43.7% phosphorus and 56.4% oxygen. What is the empirical formula of the compound?
- 4) An inorganic salt is composed of 17.6% sodium, 39.7% chromium and 42.8% oxygen. What is the empirical formula of this salt?
- Compound X contains 69.9% carbon, 6.86% hydrogen and 23.3% oxygen. Determine the empirical formula of X.
- 6) Oxalic acid has the empirical formula CHO<sub>2</sub>. Its molar mass is 90 g/mol. What is the molecular formula of oxalic acid?
- 7) The empirical formula of codeine is C<sub>18</sub>H<sub>21</sub>NO<sub>3</sub>. If the molar mass of codeine is 299 g/mol, what is its molecular formula?
- 8) A compounds molar mass is 240.28 g/mol. Its percentage composition is 75.0% carbon, 5.05% hydrogen and 20.0% oxygen. What is the compound's molecular formula?
- The wintergreen plant produces methyl salicylate, or oil of wintergreen. It can also be prepared easily in the laboratory. Methyl salicylate is 63.1% carbon, 5.31% hydrogen and 31.6% oxygen. Calculate the empirical formula of this compound.
- 10)An inorganic salt is made up of 38.8% calcium, 20.0% phosphorus, and 41.2% oxygen.
  - a) What is the empirical formula of the compound?
  - On further analysis, each formula unit of this salt is found to contain two phosphate ions.
     Predict the molecular formula of this salt.

#### Answers

- 1) Li<sub>2</sub>O
- 2) BF3
- 3) P2Os
- 4) Na<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>
- 5) C<sub>12</sub>H<sub>14</sub>O<sub>3</sub>
- 6) C<sub>2</sub>H<sub>2</sub>O<sub>4</sub>
- 7) C<sub>18</sub>H<sub>21</sub>NO<sub>3</sub>
- 8) C15H12O3
- 9) CeHeO3
- 10) a) Ca<sub>3</sub>P<sub>2</sub>O<sub>8</sub> b) Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>

# **Practice Empirical Formula Problems**

SJ Ball

#### **Practice Empirical Formula Problems:**

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