

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) _____
which cation it gains and electron it loses from it? _____ (anion/cation)
2. How placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have it. This is because it has _____ valence.
3. How _____ bond in the addition between oppositely charged ions. (ionic covalent bond) is _____
4. On a covalent bond, atoms _____, electrons, shared by all. Ionic bond, atoms _____, electrons.
5. (Electron group) Which side from left shared completely is a covalent bond. It is _____, and which group is shared equally is the _____ bond. (partial) is a covalent bond. In ionic, gain electrons, or negative charged by, losing losing from _____
6. How atoms can have up to _____ valence electrons. This usually happens in the periodic table that all have 8 valence electrons, or called _____
7. Atoms can gain, lose, or share _____ by forming ionic bonds.

B. Formulas for Ionic Compounds - Put in the right order

Write correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
COVALENT BOND				

3. Write in the chemical formula for each of the following ions: _____

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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Sebastian Brünink



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