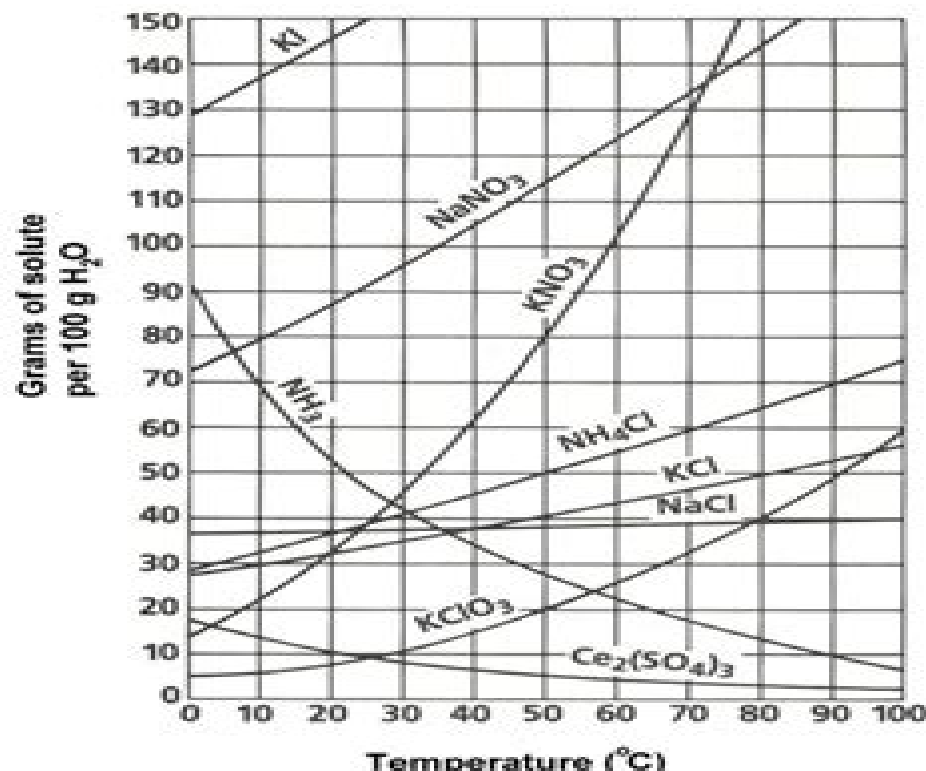


## Reading a Solubility Chart

- 1) The **curve** shows the # of grams of solute in a **saturated solution** containing 100 mL or 100 g of water at a certain temperature. (remember: 1 mL H<sub>2</sub>O = 1 g H<sub>2</sub>O)
- 2) Any amount of solute **below** the line indicates an **unsaturated solution** at a certain temperature
- 3) Any amount of solute **above** the line in which all of the solute has dissolved indicates a **supersaturated solution**.
- 4) If the amount of solute is above the line but has not all dissolved, the solution is saturated and the # grams of solute settled on the bottom of the container = total # g in solution – # g of a saturated solution at that temperature. (according to the curve)
- 5) Solutes whose curves move upward w/ increased temperature are typically solids b/c the solubility of solids increases w/ increased temperature.
- 6) Solutes whose curves move downward w/ increased temperature are typically gases b/c the solubility of gases decreases with increased temperature.



## Solubility Problems example (Many may be done by reading the solubility table)

1. At 10°C, 80 g of NaNO<sub>3</sub> will dissolve in 100 mL (a saturated solution)
2. To find the # grams needed to saturate a solution when the volume is NOT 100 mL use the following strategy to find answer: (e. g., 60 mL of H<sub>2</sub>O will be saturated with how many grams of NaNO<sub>3</sub> ?)

Start w/ known vol. x Solubility/100mL at set temp. = amount of Solute needed to saturate

$$\text{Ex. } 60 \text{ mL H}_2\text{O} \times \frac{80 \text{ g NaNO}_3}{100 \text{ mL H}_2\text{O}} = 48 \text{ g NaNO}_3 \text{ needed to saturate solution}$$

or if the chart is in units of 100 g of H<sub>2</sub>O use the density of water conversion 1mL H<sub>2</sub>O= 1 g H<sub>2</sub>O

$$\text{Ex. } 60 \text{ mL H}_2\text{O} \times \frac{1 \text{ g H}_2\text{O}}{1 \text{ mL H}_2\text{O}} \times \frac{80 \text{ g NaNO}_3}{100 \text{ g H}_2\text{O}} = 48 \text{ g NaNO}_3$$

# Reading The Solubility Chart Problems Answer Key

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## **Table of Contents Reading The Solubility Chart Problems Answer Key**

1. Understanding the eBook Reading The Solubility Chart Problems Answer Key
  - The Rise of Digital Reading Reading The Solubility Chart Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Reading The Solubility Chart Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reading The Solubility Chart Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Reading The Solubility Chart Problems Answer Key
  - Personalized Recommendations
  - Reading The Solubility Chart Problems Answer Key User Reviews and Ratings
  - Reading The Solubility Chart Problems Answer Key and Bestseller Lists
5. Accessing Reading The Solubility Chart Problems Answer Key Free and Paid eBooks

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- Reading The Solubility Chart Problems Answer Key eBook Subscription Services
- Reading The Solubility Chart Problems Answer Key Budget-Friendly Options
- 6. Navigating Reading The Solubility Chart Problems Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Reading The Solubility Chart Problems Answer Key Compatibility with Devices
  - Reading The Solubility Chart Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reading The Solubility Chart Problems Answer Key
  - Highlighting and Note-Taking Reading The Solubility Chart Problems Answer Key
  - Interactive Elements Reading The Solubility Chart Problems Answer Key
- 8. Staying Engaged with Reading The Solubility Chart Problems Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Reading The Solubility Chart Problems Answer Key
- 9. Balancing eBooks and Physical Books Reading The Solubility Chart Problems Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reading The Solubility Chart Problems Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Reading The Solubility Chart Problems Answer Key
  - Setting Reading Goals Reading The Solubility Chart Problems Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reading The Solubility Chart Problems Answer Key
  - Fact-Checking eBook Content of Reading The Solubility Chart Problems Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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