

2. Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.

a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)



b. What type of reaction is this? single replacement

c. If 16.2 g of Mg are heated with 12.0 g of H_2O , which is the limiting reactant?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.0 \text{ g H}_2\text{O}$$

There is less H_2O (12.0 g) in our initial amount so
 H_2O is the L.R

d. How many moles of the excess reactant remain after the reaction is complete?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} = 1.3 \text{ mol H}_2\text{O}$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.666 \text{ mol H}_2\text{O}$$

$$1.3 \text{ mol H}_2\text{O} - 0.666 \text{ mol H}_2\text{O} = 0.634 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Mg}}{2 \text{ mol H}_2\text{O}} = \boxed{0.317 \text{ mol Mg}}$$

e. How many grams of each product are formed?

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol Mg}(\text{OH})_2}{2 \text{ mol H}_2\text{O}} \times \frac{58.32 \text{ g Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{OH})_2} = 19.4 \text{ g Mg}(\text{OH})_2$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} \times \frac{2.02 \text{ g H}_2}{1 \text{ mol H}_2} = 6.28 \text{ g H}_2$$

e. You perform this reaction, and obtain 16.0 g $\text{Mg}(\text{OH})_2$. What is your percent yield?

$$\% \text{ yield} = \frac{\text{actual} \times 100}{\text{theoretical}}$$

$$\frac{16.0 \text{ g Mg}(\text{OH})_2}{19.4 \text{ g Mg}(\text{OH})_2} \times 100 = 82.5\%$$

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 - The Rise of Digital Reading Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Hall Chemistry Worksheets 7 2 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Personalized Recommendations
 - Practice Hall Chemistry Worksheets 7 2 Answer Key User Reviews and Ratings
 - Practice Hall Chemistry Worksheets 7 2 Answer Key and Bestseller Lists
5. Accessing Practice Hall Chemistry Worksheets 7 2 Answer Key Free and Paid eBooks
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- Practice Hall Chemistry Worksheets 7 2 Answer Key Budget-Friendly Options
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 - Practice Hall Chemistry Worksheets 7 2 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Highlighting and Note-Taking Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Interactive Elements Practice Hall Chemistry Worksheets 7 2 Answer Key
- 8. Staying Engaged with Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Hall Chemistry Worksheets 7 2 Answer Key
- 9. Balancing eBooks and Physical Books Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Hall Chemistry Worksheets 7 2 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Setting Reading Goals Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Fact-Checking eBook Content of Practice Hall Chemistry Worksheets 7 2 Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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