

Geometry Notes

6-3 Conditions for Parallelograms

In Section 6-2, we looked at the definition of a parallelogram and special properties of parallelograms. In this section, we will work backwards. Given a quadrilateral, how do we know whether or not it is a parallelogram? There are several ways to tell.

Conditions for Parallelograms

A quadrilateral is a parallelogram if:

- 1) **Both** pairs of opposite sides are **parallel** (remember this is the definition.)
- 2) **Both** pairs of opposite **sides** are **congruent**.
- 3) **Both** pairs of opposite **angles** are **congruent**.
- 4) **One** pair of opposite **sides** is both **parallel AND congruent**.
- 5) An angle is **supplementary** to **both** of its **consecutive angles**.
- 6) **Both diagonals bisect each other**.

Ex. 1: Determine if each of the following must be a parallelogram. If so, give the reason.

a)



Yes, one pair of sides parallel and $\cong$

b)



Yes, one angle is supp. to both consecutive angles

c)



No

d)



Technically, yes. You can show that the parallel sides are $\cong$ by ASA.

Practice A Conditions For Parallelograms

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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note

that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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