



NATIONAL  
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2  
(CHEMISTRY)  
EXEMPLAR

MARKS: 100  
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.  
Use permanent marker for writing and drawing.

## QUESTION 3

### PHYSICAL SCIENCES PAPER 2

### INTERMOLECULAR FORCES

**GRADE 11**

**NOVEMBER 2020**

Examiner's use: \_\_\_\_\_ Physical Sciences P2

QUESTION 3 starts on a NEW page.

- 3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the  $\Delta H_{\text{vap}}$  and  $\Delta H_{\text{fusion}}$  for 1-hexanol in kJ mol<sup>-1</sup> and 1 mol of 1-hexanol.

Compound	Molecular mass (M <sub>r</sub> )	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol composed of one (A) or two (B) and list one of its intermolecular forces? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What (one) of the compounds (A or B) has a (threefold) greater mass (molar mass)? (2)
- 3.1.5 Which compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Which compound (A or B) will have the higher surface tension? (2)

Compound B (M<sub>r</sub> 102) is shown in the table. Use a suitable intermolecular force that both compounds A and B have to identify liquid boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound B is (threefold) higher than compound A and B is (threefold) higher than 1-hexanol in terms of intermolecular forces. (2)

- 3.2 Substances C1 is added to water whereas 1-hexanol (substance C2) are added to hexane. See also the diagram below.



- 3.2.1 Is substance C1 or C2, oil or water? (2)
- 3.2.2 Explain the process in QUESTION 3.1.1 about the boiling in the intermolecular forces involved. (2)

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# Physical Science Paper2 November 2014

**David Kirk**



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