

## Periodic Trends Worksheet

Directions: Use your class materials to complete the following questions.

1. Write the following elements by increasing atomic radius: sodium, calcium, oxygen, phosphorus.

**Na, Ca, P, O**

2. Write the following elements by increasing electronegativity: sodium, oxygen, calcium, aluminum.

**Na, Al, Ca, O**

3. Which atom has the higher first ionization energy: Na or Ca?

**Ca has a higher first ionization energy than Na because it has more protons.**

4. Which two elements in the same family generally have similar properties?

**There have been two elements that share a same periodic table group.**

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

**Period 7, francium, cesium, barium, potassium, rubidium, and sodium.**

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

**Period 1, hydrogen, helium, lithium, beryllium, boron, carbon, nitrogen, oxygen, and fluorine.**

8. Write the atoms in each group that have the largest atomic radius.

Li	Na	K
Rb	Cs	Fr
Be	Mg	Ca
Sc	Y	La
Ti	Zr	Hf
V	Nb	Ta
Cr	Mo	W
Mn	Tc	Re
Fe	Ru	Rh
Co	Pd	Pt
Ni	Cu	Ag
Cd	Hg	

Al	B	Ga
In	Tl	
Sn	Pb	
Sb	Bi	
Te	Po	
As	Sb	
Se	Te	
Br	I	At
Kr	Xe	Rn

9. Circle the atoms in each group that have the greatest electronegativity.

# Physical Science If8767 Pg 63 Key

**Ensheng Dong**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Physical Science If8767 Pg 63 Key**

1. Understanding the eBook Physical Science If8767 Pg 63 Key
  - The Rise of Digital Reading Physical Science If8767 Pg 63 Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Pg 63 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science If8767 Pg 63 Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Pg 63 Key
  - Personalized Recommendations
  - Physical Science If8767 Pg 63 Key User Reviews and Ratings
  - Physical Science If8767 Pg 63 Key and Bestseller Lists
5. Accessing Physical Science If8767 Pg 63 Key Free and Paid eBooks
  - Physical Science If8767 Pg 63 Key Public Domain eBooks
  - Physical Science If8767 Pg 63 Key eBook Subscription Services
  - Physical Science If8767 Pg 63 Key Budget-Friendly Options
6. Navigating Physical Science If8767 Pg 63 Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science If8767 Pg 63 Key Compatibility with Devices
  - Physical Science If8767 Pg 63 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science If8767 Pg 63 Key
  - Highlighting and Note-Taking Physical Science If8767 Pg 63 Key
  - Interactive Elements Physical Science If8767 Pg 63 Key
8. Staying Engaged with Physical Science If8767 Pg 63 Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Physical Science If8767 Pg 63 Key
- 9. Balancing eBooks and Physical Books Physical Science If8767 Pg 63 Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science If8767 Pg 63 Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science If8767 Pg 63 Key
  - Setting Reading Goals Physical Science If8767 Pg 63 Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science If8767 Pg 63 Key
  - Fact-Checking eBook Content of Physical Science If8767 Pg 63 Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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