

Name \_\_\_\_\_

# Physical Science Chemical Bonding Study Guide

## A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? \_\_\_\_\_ (anion)  
Which cation is always a transition metal? \_\_\_\_\_ (cation)
2. An element in the halogen group has \_\_\_\_\_ valence electrons, and tends to \_\_\_\_\_ 1 electron to have 8. This is because it has \_\_\_\_\_ valence.
3. An \_\_\_\_\_ bond is the attraction between oppositely charged ions. HCl is ionic and is \_\_\_\_\_.
4. An ionic bond is formed between \_\_\_\_\_, which is attracted by an ionic bond, and \_\_\_\_\_, which is \_\_\_\_\_.
5. Elements in the \_\_\_\_\_ group always bond with a charge of \_\_\_\_\_, and elements in the \_\_\_\_\_ group always bond with a charge of \_\_\_\_\_. This is because \_\_\_\_\_ is \_\_\_\_\_, so \_\_\_\_\_ is \_\_\_\_\_.
6. An ionic bond is formed between \_\_\_\_\_, which is \_\_\_\_\_, and \_\_\_\_\_, which is \_\_\_\_\_.
7. Elements in the \_\_\_\_\_ group have \_\_\_\_\_ valence electrons, and tend to \_\_\_\_\_ 1 electron to have 8.

## B. Formulas for Ionic Compounds - Fill in the blank below

Write the correct chemical formula for the compound.	Aluminum	Barium	Phosphorus	Chlorine
_____	_____	_____	_____	_____

## C. Write the chemical formula for each of the following ions:

Write your answer below the Lewis structure, and use arrows to show the transfer of electrons. Use the periodic table to find the oxidation number. They will combine so that the oxidation numbers cancel out to zero. Use the criss cross method if the oxidation numbers do not cancel.

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**Minjie Lin**



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