

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, Cl, O

3. Which atom has the higher electronegativity: sodium or sulfur?

Sulfur (sulfur is to the right of sodium and closer to fluorine)

4. Which are stronger in the same family: generally fluorine or chlorine?

Fluorine (fluorine is above chlorine and closer to fluorine)

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7, francium (francium is at the bottom left corner)

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Period 1, hydrogen (hydrogen is at the top left corner)

8. Write the atoms in each group that have the largest atomic radius.

Li	Be
Na	Mg
K	Ca
Rb	Sr
Cs	Ba
Fr	Ra

Al	Si	P	S
Ga	Ge	As	Se
In	Sn	Sb	Te
Tl	Pb	Bi	Po

9. Circle the atom in each group that has the greatest electronegativity.

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Table of Contents Physical Science If8767 Pg 63 Key

1. Understanding the eBook Physical Science If8767 Pg 63 Key
 - The Rise of Digital Reading Physical Science If8767 Pg 63 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Pg 63 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Physical Science If8767 Pg 63 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Pg 63 Key
 - Personalized Recommendations

- Physical Science If8767 Pg 63 Key User Reviews and Ratings
- Physical Science If8767 Pg 63 Key and Bestseller Lists
- 5. Accessing Physical Science If8767 Pg 63 Key Free and Paid eBooks
 - Physical Science If8767 Pg 63 Key Public Domain eBooks
 - Physical Science If8767 Pg 63 Key eBook Subscription Services
 - Physical Science If8767 Pg 63 Key Budget-Friendly Options
- 6. Navigating Physical Science If8767 Pg 63 Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science If8767 Pg 63 Key Compatibility with Devices
 - Physical Science If8767 Pg 63 Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science If8767 Pg 63 Key
 - Highlighting and Note-Taking Physical Science If8767 Pg 63 Key
 - Interactive Elements Physical Science If8767 Pg 63 Key
- 8. Staying Engaged with Physical Science If8767 Pg 63 Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science If8767 Pg 63 Key
- 9. Balancing eBooks and Physical Books Physical Science If8767 Pg 63 Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science If8767 Pg 63 Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science If8767 Pg 63 Key
 - Setting Reading Goals Physical Science If8767 Pg 63 Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science If8767 Pg 63 Key
 - Fact-Checking eBook Content of Physical Science If8767 Pg 63 Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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