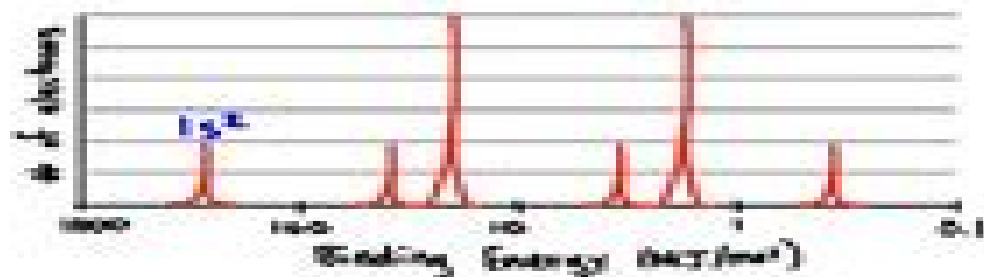


Reading PES Practice

Name: _____

Spectrum 1:



1. Label the electron configuration for each peak above that specific peak. The first one was done for you.

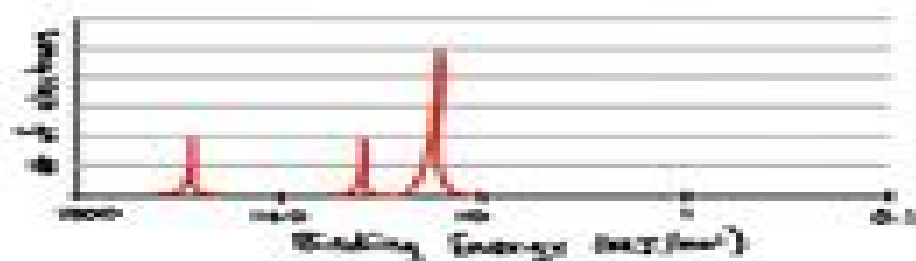
2. This element has

_____ core electrons,

_____ valence electrons.

- The element is _____. It forms _____ (charge sign and value) ions.
- On the spectrum, put a star above the peak(s) whose electrons are furthest from the nucleus.
- On the spectrum, put an X above the peak(s) whose electrons are in a p sublevel.
- This element bonds with oxygen. Write the chemical formula for the resultant compound **AND** select another element that would bond with oxygen in the same proportion.

Spectrum 2:



1. Label the electron configuration for each peak above that specific peak.

2. This element has

_____ core electrons,

_____ valence electrons.

- The element is _____. It forms _____ (charge and numerical value) ions.
- On the spectrum, put a star above the peak(s) whose electrons are hardest to remove from the atom.

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Neil D. Jespersen, Pamela Kerrigan



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