

2. Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.

a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)



b. What type of reaction is this? single replacement

c. If 16.2 g of Mg are heated with 12.0 g of  $\text{H}_2\text{O}$ , which is the limiting reactant?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.0 \text{ g H}_2\text{O}$$

There is less  $\text{H}_2\text{O}$  (12.0 g) in our initial amount so  
 $\text{H}_2\text{O}$  is the L.R

d. How many moles of the excess reactant remain after the reaction is complete?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} = 1.3 \text{ mol H}_2\text{O}$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.666 \text{ mol H}_2\text{O}$$

$$1.3 \text{ mol H}_2\text{O} - 0.666 \text{ mol H}_2\text{O} = 0.634 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Mg}}{2 \text{ mol H}_2\text{O}} = \boxed{0.317 \text{ mol Mg}}$$

e. How many grams of each product are formed?

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol Mg}(\text{OH})_2}{2 \text{ mol H}_2\text{O}} \times \frac{58.32 \text{ g Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{OH})_2} = 19.4 \text{ g Mg}(\text{OH})_2$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} \times \frac{2.02 \text{ g H}_2}{1 \text{ mol H}_2} = 6.28 \text{ g H}_2$$

e. You perform this reaction, and obtain 16.0 g  $\text{Mg}(\text{OH})_2$ . What is your percent yield?

$$\% \text{ yield} = \frac{\text{actual} \times 100}{\text{theoretical}}$$

$$\frac{16.0 \text{ g Mg}(\text{OH})_2}{19.4 \text{ g Mg}(\text{OH})_2} \times 100 = 82.5\%$$

# Sample Problems H Gas Stoichiometry Answers

**Ying Liu**



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## **Table of Contents Sample Problems H Gas Stoichiometry Answers**

1. Understanding the eBook Sample Problems H Gas Stoichiometry Answers
  - The Rise of Digital Reading Sample Problems H Gas Stoichiometry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Sample Problems H Gas Stoichiometry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Sample Problems H Gas Stoichiometry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Sample Problems H Gas Stoichiometry Answers
  - Personalized Recommendations
  - Sample Problems H Gas Stoichiometry Answers User Reviews and Ratings
  - Sample Problems H Gas Stoichiometry Answers and Bestseller Lists
5. Accessing Sample Problems H Gas Stoichiometry Answers Free and Paid eBooks
  - Sample Problems H Gas Stoichiometry Answers Public Domain eBooks
  - Sample Problems H Gas Stoichiometry Answers eBook Subscription Services
  - Sample Problems H Gas Stoichiometry Answers Budget-Friendly Options
6. Navigating Sample Problems H Gas Stoichiometry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Sample Problems H Gas Stoichiometry Answers Compatibility with Devices
  - Sample Problems H Gas Stoichiometry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Sample Problems H Gas Stoichiometry Answers
  - Highlighting and Note-Taking Sample Problems H Gas Stoichiometry Answers
  - Interactive Elements Sample Problems H Gas Stoichiometry Answers

8. Staying Engaged with Sample Problems H Gas Stoichiometry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Sample Problems H Gas Stoichiometry Answers
9. Balancing eBooks and Physical Books Sample Problems H Gas Stoichiometry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Sample Problems H Gas Stoichiometry Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Sample Problems H Gas Stoichiometry Answers
  - Setting Reading Goals Sample Problems H Gas Stoichiometry Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Sample Problems H Gas Stoichiometry Answers
  - Fact-Checking eBook Content of Sample Problems H Gas Stoichiometry Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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