

Name _____ Date _____

DOSAGE CALCULATION

CONVERSIONS

1 mg = 1000 mcg

1 oz = 2 Tbsp

3 Tsp = 1 Tbsp

1 L = 1000 mL

1 kg = 1000 g

30 mL = 1 oz

5 mL = 1 Tsp

1 gr = 1000 mcg

8 oz = 1 cup

15 mL = 1 Tbsp

1 mL = 1 cc

1 kg = 2,2 lbs

1. MD orders Morphine 0.5 mg IV every 4 hours as needed for pain. The vial is labeled 2 mg/mL. How many mL will you give per dose?
2. MD orders Benadryl 25 mg IV as needed for itching. The vial is labeled 1000 mcg/mL. How many mL will you give per dose?
3. MD orders Digoxin 0.125 mg IV stat. The vial is labeled 500 mcg/2mL. How many mL will you give per dose?
4. MD orders Versed 4 mg IV pre-procedural. The vial is labeled 1 mg/mL. How many mL will you give per dose?
5. MD orders Heparin 100 units subq daily. The vial is labeled 50 units/mL. How many mL will you give per dose?

Sample Medical Math Problems

L Manion



Sample Medical Math Problems:

This book delves into Sample Medical Math Problems. Sample Medical Math Problems is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Sample Medical Math Problems, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Sample Medical Math Problems
- Chapter 2: Essential Elements of Sample Medical Math Problems
- Chapter 3: Sample Medical Math Problems in Everyday Life
- Chapter 4: Sample Medical Math Problems in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Sample Medical Math Problems. This chapter will explore what Sample Medical Math Problems is, why Sample Medical Math Problems is vital, and how to effectively learn about Sample Medical Math Problems.

3. In chapter 2, the author will delve into the foundational concepts of Sample Medical Math Problems. The second chapter will elucidate the essential principles that must be understood to grasp Sample Medical Math Problems in its entirety.

4. In chapter 3, this book will examine the practical applications of Sample Medical Math Problems in daily life. This chapter will showcase real-world examples of how Sample Medical Math Problems can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Sample Medical Math Problems in specific contexts. The fourth chapter will explore how Sample Medical Math Problems is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Sample Medical Math Problems. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Sample Medical Math Problems.

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add H^+ ions or the H_2O pair to fully balance the equation in this video we'll walk through this process for the reaction between dichromate $\text{Cr}_2\text{O}_7^{2-}$ and chloride Cl^- ions in acidic solution created by jay

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web balance each half reaction in basic solution $4\text{Fe}^{2+} + 5\text{H}_2\text{O} + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 4\text{Fe}^{3+} + 2\text{NO}_3^- + 2\text{MnO}_2 + 3\text{Cr}_2\text{O}_3 + 3\text{SO}_2 + 2\text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O} + 2\text{Cr}^{3+} + 9\text{H}_2\text{O}$ $2\text{N}_2\text{O}_4 + \text{O}_2 \rightarrow 2\text{Cr}^{3+} + 10\text{H}_2\text{O} + 4\text{H}^+$ $11\text{PbO}_2 + \text{IO}_3^- \rightarrow \text{IO}_3^- + 2\text{H}_2\text{SO}_4$

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the

reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h⁺ o²⁻ h⁺ and e⁻ in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and