

Practice Final Exam Chem 112

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ____ 1. Three hundred milliliters (150 mL) of 4.50 M sugar solution is diluted to 600 mL by the addition of pure water. Which of these statements about the solution is correct?
a. The moles of water remain the same. b. The moles of sugar decrease. c. The concentration of sugar increases to 5.0 M. d. The moles of sugar remain constant. e. All of these statements are incorrect.
- ____ 2. How many moles of solute are present in 75.0 grams of solution that is 1.30% by mass glucose ($C_6H_{12}O_6$)? $MM(C_6H_{12}O_6) = 180.16 \text{ g/mol}$
a. 5.4×10^{-3} moles b. 0.800 moles c. 176 moles d. 1.76×10^{-3} moles e. 225 moles
- ____ 3. A concentrated reagent is sold as a solution that is 60.0 mass % solute. If the molar mass of the solute is 244 g/mol and the density of this solution is 1.20 g/mL, what is the molarity of the solution?
a. 10.0 M b. 2.95 M c. 15.5 M d. 0.72 M e. 1.45 M
- ____ 4. The boiling point of pure benzene (C_6H_6) is 80.10 °C. A solution of 14.2 g of a solid dissolved in 64.0 g of benzene boils at 81.09 °C. If k_b for benzene is 2.53 °C/molal, the molar mass of the solid is:
a. 662 g/mol b. 56.0 g/mol c. 92.0 g/mol d. 628 g/mol e. 568 g/mol
- ____ 5. What is the freezing point for a 0.71 m solution of glycerol ($C_3H_8O_3$) in ethanol? Freezing point(ethanol) = -114.6 °C; $K_f(\text{ethanol}) = 1.99 \text{ °C/m}$
a. 0.10 °C b. -113.2 °C c. -1.41 °C d. -116.0 °C e. none of these
- ____ 6. At 450 °C, $K_c = 0.231$ for the reaction:



What is the value of the equilibrium constant for the following reaction at 450 °C?



- a. 0.481 b. 0.156 c. 0.231 d. 0.053 e. 4.329

- ____ 7. Consider the following system at equilibrium and its corresponding concentration equilibrium constant:



What is the value of K_p for this reaction at 470 °C?

- a. $K_p = 3.80 \times 10^{-6}$ b. $K_p = 8.53 \times 10^{-6}$ c. $K_p = 4.32 \times 10^{-5}$ d. $K_p = 1.10 \times 10^{-4}$ e. $K_p = 1.19 \times 10^4$

- ____ 8. For the reaction $\text{I}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{ICl}(\text{g})$, the relation between K_p and K_c is:
a. $K_p = K_c$ b. $K_p = K_c(RT)^2$ c. $K_p = K_c(RT)^{-2}$ d. $K_c = K_p(RT)^2$ e. $K_p = K_c(RT)^4$

Practice General Chemistry 112 Final Exam

Ensheng Dong



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Table of Contents Practice General Chemistry 112 Final Exam

1. Understanding the eBook Practice General Chemistry 112 Final Exam
 - The Rise of Digital Reading Practice General Chemistry 112 Final Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice General Chemistry 112 Final Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice General Chemistry 112 Final Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice General Chemistry 112 Final Exam
 - Personalized Recommendations
 - Practice General Chemistry 112 Final Exam User Reviews and Ratings
 - Practice General Chemistry 112 Final Exam and Bestseller Lists

5. Accessing Practice General Chemistry 112 Final Exam Free and Paid eBooks
 - Practice General Chemistry 112 Final Exam Public Domain eBooks
 - Practice General Chemistry 112 Final Exam eBook Subscription Services
 - Practice General Chemistry 112 Final Exam Budget-Friendly Options
6. Navigating Practice General Chemistry 112 Final Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice General Chemistry 112 Final Exam Compatibility with Devices
 - Practice General Chemistry 112 Final Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice General Chemistry 112 Final Exam
 - Highlighting and Note-Taking Practice General Chemistry 112 Final Exam
 - Interactive Elements Practice General Chemistry 112 Final Exam
8. Staying Engaged with Practice General Chemistry 112 Final Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice General Chemistry 112 Final Exam
9. Balancing eBooks and Physical Books Practice General Chemistry 112 Final Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice General Chemistry 112 Final Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Practice General Chemistry 112 Final Exam
 - Setting Reading Goals Practice General Chemistry 112 Final Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Practice General Chemistry 112 Final Exam
 - Fact-Checking eBook Content of Practice General Chemistry 112 Final Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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