

QUESTION 2

- 2.1 The graph of molecular size versus the boiling point is given below.
The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

| NAME OF SUBSTANCE | FORMULA | MOLACULAR MASS (g) | BOILING POINT ($^{\circ}\text{C}$) |
|-------------------|---------------|--------------------|--------------------------------------|
| Ammonia | NH_3 | 17 | - 33 |
| Phosphine | PH_3 | x | - 87,4 |

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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 - The Rise of Digital Reading Physical Science Department Grade Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Department Grade Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Department Grade Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Department Grade Question Paper
 - Personalized Recommendations
 - Physical Science Department Grade Question Paper User Reviews and Ratings
 - Physical Science Department Grade Question Paper and Bestseller Lists
5. Accessing Physical Science Department Grade Question Paper Free and Paid eBooks

- Physical Science Department Grade Question Paper Public Domain eBooks
 - Physical Science Department Grade Question Paper eBook Subscription Services
 - Physical Science Department Grade Question Paper Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Physical Science Department Grade Question Paper Compatibility with Devices
 - Physical Science Department Grade Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Physical Science Department Grade Question Paper
 - Interactive Elements Physical Science Department Grade Question Paper
8. Staying Engaged with Physical Science Department Grade Question Paper
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Department Grade Question Paper
9. Balancing eBooks and Physical Books Physical Science Department Grade Question Paper
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Department Grade Question Paper
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Department Grade Question Paper
- Setting Reading Goals Physical Science Department Grade Question Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Department Grade Question Paper
- Fact-Checking eBook Content of Physical Science Department Grade Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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