

PHYSICAL SCIENCES - GRADE 10

1 Revision of matter (Grade 9)

1.1 Matter and classification

- Matter is made up of particles (atoms or molecules) and it is the properties of the atoms or molecules that determine the characteristics and reactivity of that matter.
- The properties of matter include: strength; density; melting and boiling points; whether the material is brittle, malleable or ductile; whether it is magnetic or not; and its electrical and thermal conductivity.
- Elements and compounds are classified as pure because they contain particles that are all the same. Substances can be classified as pure – elements and compounds
 - or mixtures. Pure substances contain particles that are all the same. Elements and compounds can be represented by symbols and formulae. Some of the names, symbols or formulae of a number of common elements and compounds that you should have learnt are the following:
 - Elements: S for sulphur; C for carbon; P for phosphorus (note the spelling); H for hydrogen; O for oxygen; He for helium; N for nitrogen; F for fluorine; Mg for magnesium; Ca for calcium; Zn for zinc.
 - Compounds: H_2O for water; H_2SO_4 for sulphuric acid; HCl for hydrochloric acid; NaCl for sodium chloride (table salt); $CaCO_3$ for calcium carbonate; KNO_3 for potassium nitrate; Na_3PO_4 for sodium phosphate.
- You will notice that the naming of compounds follows these rules
 - The metal or hydrogen is always named first, then the non-metal (potassium iodide: KI; hydrogen chloride: HCl (also called hydrochloric acid)).
 - The name of a non-metal ion of an element always ends in -ide (for instance, sodium chloride: NaCl).
 - The name of a non-metal complex ion (more than one element in the ion) usually ends in -ate or -ite (for instance, copper sulphate: $CuSO_4$; potassium sulphite: K_2SO_3). An exception to this is hydroxide (for instance, sodium hydroxide: NaOH).
- Mixtures consist of different substances, and can be homogeneous (have the same 'look' throughout) or heterogeneous (you can see the different substances). They can be separated in various ways, for instance filtering, sieving, chromatography, distilling or dissolving one constituent.
- Some of the properties of metals are that they conduct electricity and heat, and they are malleable and ductile.
- Non-metals are non-conductors of electricity and heat. They are gases, liquids or brittle solids, and bond with each other.
- Metalloids are elements on the border between metals and non-metals, and they have both metal and non-metal properties.

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Wenbin Ji



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Table of Contents Physical Science Grade 1study Notes Third Term

1. Understanding the eBook Physical Science Grade 1study Notes Third Term
 - The Rise of Digital Reading Physical Science Grade 1study Notes Third Term
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 1study Notes Third Term
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grade 1study Notes Third Term
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 1study Notes Third Term
 - Personalized Recommendations
 - Physical Science Grade 1study Notes Third Term User Reviews and Ratings
 - Physical Science Grade 1study Notes Third Term and Bestseller Lists

5. Accessing Physical Science Grade 1study Notes Third Term Free and Paid eBooks
 - Physical Science Grade 1study Notes Third Term Public Domain eBooks
 - Physical Science Grade 1study Notes Third Term eBook Subscription Services
 - Physical Science Grade 1study Notes Third Term Budget-Friendly Options
6. Navigating Physical Science Grade 1study Notes Third Term eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Grade 1study Notes Third Term Compatibility with Devices
 - Physical Science Grade 1study Notes Third Term Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Grade 1study Notes Third Term
 - Highlighting and Note-Taking Physical Science Grade 1study Notes Third Term
 - Interactive Elements Physical Science Grade 1study Notes Third Term
8. Staying Engaged with Physical Science Grade 1study Notes Third Term
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade 1study Notes Third Term
9. Balancing eBooks and Physical Books Physical Science Grade 1study Notes Third Term
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grade 1study Notes Third Term
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grade 1study Notes Third Term
 - Setting Reading Goals Physical Science Grade 1study Notes Third Term
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Grade 1study Notes Third Term
 - Fact-Checking eBook Content of Physical Science Grade 1study Notes Third Term
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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