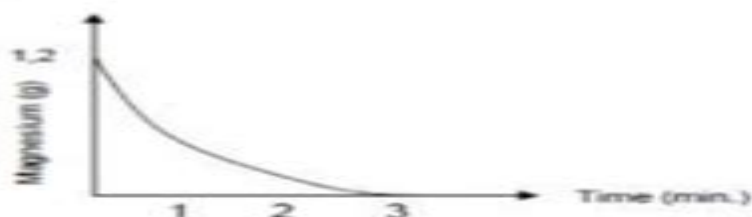


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records the change in the mass of magnesium as the reaction proceeds at regular intervals. The balanced chemical equation for the reaction is:



The change in the mass of magnesium during the reaction is shown on the graph below.



6.2.1 Identify the limiting agent in this reaction. Give a reason for the answer. (2)

6.2.2 Calculate the number of moles of unreacted hydrochloric acid in the test tube after 3 minutes. (7)

[19]

Hide Answers

QUESTION/VRAAG 6

6.1.1 $2\text{H}_2\text{O}_2\text{(aq)} \checkmark \rightarrow 2\text{H}_2\text{O(l)} + \text{O}_2\text{(g)} \checkmark$ (2)

6.1.2 The catalyst lowers the activation energy of the reaction ✓✓

Accept: catalyst speeds up the reaction

In Katalisator verlaag die aktiveringsenergie van die reaksie ✓✓

Aanvaar: katalisator laat die reaksie vinniger plaasvind (2)

6.1.3 **OPTION 1/OPSIE 1**

$$n = \frac{V}{V_m} \checkmark$$

$$n = \frac{0,6}{24,45} \checkmark$$

$$n = 0,0245 \text{ mole/mol O}_2 \text{ produced/gevorm}$$

$$\text{H}_2\text{O}_2 : \text{O}_2 \checkmark$$

$$2 : 1 \checkmark$$

$$n = 0,049 \text{ mole/mol H}_2\text{O}_2 \text{ reacted/reageer}$$

$$n = \frac{m}{M} \checkmark$$

$$0,049 = \frac{m}{34} \checkmark$$

OPTION 2/OPSIE 2

From the balanced equation:

Vanaf gebalanseerde vergelyking:



$$X = \frac{68 \times 0,6}{24,45} \checkmark$$

$$X = 1,67 \text{ g} \checkmark$$

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1. Understanding the eBook Physical Science Exemplar November 2014 Grade 11
 - The Rise of Digital Reading Physical Science Exemplar November 2014 Grade 11
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Exemplar November 2014 Grade 11
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Exemplar November 2014 Grade 11
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Exemplar November 2014 Grade 11

- Personalized Recommendations
 - Physical Science Exemplar November 2014 Grade 11 User Reviews and Ratings
 - Physical Science Exemplar November 2014 Grade 11 and Bestseller Lists
5. Accessing Physical Science Exemplar November 2014 Grade 11 Free and Paid eBooks
 - Physical Science Exemplar November 2014 Grade 11 Public Domain eBooks
 - Physical Science Exemplar November 2014 Grade 11 eBook Subscription Services
 - Physical Science Exemplar November 2014 Grade 11 Budget-Friendly Options
 6. Navigating Physical Science Exemplar November 2014 Grade 11 eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Exemplar November 2014 Grade 11 Compatibility with Devices
 - Physical Science Exemplar November 2014 Grade 11 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Exemplar November 2014 Grade 11
 - Highlighting and Note-Taking Physical Science Exemplar November 2014 Grade 11
 - Interactive Elements Physical Science Exemplar November 2014 Grade 11
 8. Staying Engaged with Physical Science Exemplar November 2014 Grade 11
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Exemplar November 2014 Grade 11
 9. Balancing eBooks and Physical Books Physical Science Exemplar November 2014 Grade 11
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Exemplar November 2014 Grade 11
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Science Exemplar November 2014 Grade 11
 - Setting Reading Goals Physical Science Exemplar November 2014 Grade 11
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Science Exemplar November 2014 Grade 11

- Fact-Checking eBook Content of Physical Science Exemplar November 2014 Grade 11
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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therapieziele 108 6 3 empfehlungen zur therapie nach schlaganfall 110 6 3 1 bewegung und mobilität 110 arm und
beinparesen 110 orofaziale paresen 111 apraxie 112 spastik 112 störungen der posturalen kontrolle 114

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