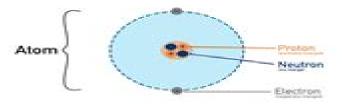
### C11- PHYSICAL SCIENCE

Third Monthly Test Reviewer

Coverage: MODULE 1 UNIT 1: Atomic Theory UNIT 2: The Discovery and Synthesis of Chemical Elements

#### **UNIT 1: Atomic Theory**

 Atoms are composed of protons, neutrons, and electrons.



**Table 2.3 Properties of Subatomic Particles** 

Particle	Symbol	Meso (kg)	Mass (pro-	Relative Charge
Prophers	800	2.679 × 207	1	
Pleasablean	407	1.475 × 10°	1	- 0
Electronic .		4.506 × 50°	0.00055	-8

- An atom is mostly empty space with a highly dense, positive center called nucleus.
- For a neutral atom, the number of protons is equivalent to the number of electrons.
   (Lithium has 3 protons therefore it

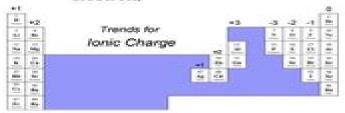
(Lithium has 3 protons therefore it also has 3 electrons)

- The atomic number of an element is also equivalent to its number of protons.
   (Nitrogen has an atomic number of 7 therefore it has 7 protons in its nucleus)
- The atomic mass (mass number) is primarily determined by the number of protons and neutrons.

- Isotopes are atoms with the same atomic number but different mass numbers. They differ in the number of neutrons.
   (Carbon-12, Carbon-13, Carbon-14)
- Ion is an element or atom that has a positive or negative net charge.
   It is charged because the number of electrons does not equal the number of protons in the atom.

## Two types:

- Cation- positive ion (loses electron)
- Anion- negative ion (gains electron)



- Take note that when an atom is not electrically neutral, it means that it does not have the same number of protons and electrons.
- When an atom or element is positively charged, it means that it has more protons in its nucleus.

## Historical Development of Atom

- Democritus & Leucippus: Atomism: All matter is made up of very small, indivisible particles called atoms (atomos).
- Aristotle: Four elements in nature (air, fire, water, and earth)
- John Dalton: Billiard Ball Model/Solid Sphere Model: Atomic Theory: All matter is made up of very small, indivisible particles called atoms.

# **Phisical Science Gradeexempler Caps 2014**

**Karin Nielsen-Saines** 

Phisical Science Gradeexempler Caps 2014:

The book delves into Phisical Science Gradeexempler Caps 2014. Phisical Science Gradeexempler Caps 2014 is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Phisical Science Gradeexempler Caps 2014, encompassing both the fundamentals and more intricate discussions.

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  - Chapter 3: Phisical Science Gradeexempler Caps 2014 in Everyday Life
  - Chapter 4: Phisical Science Gradeexempler Caps 2014 in Specific Contexts
  - ∘ Chapter 5: Conclusion
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# **Phisical Science Gradeexempler Caps 2014 Introduction**

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