

3. Refer to the data table in Question 2.

a. How do the reduced ratios in the last column compare to the coefficients in the reaction shown in Model 1?

*The ratios are the same as the coefficients in Model 1: 1:3:2.*

b. Use mathematical concepts to explain how your answer in part a is possible.

*The coefficients were all multiplied by the same number, so they all ended back to the same thing.*

4. Even 538 is a small number of molecules to use in a reaction. Typically chemists use much larger numbers of molecules. (Recall that one mole is equal to  $6.02 \times 10^{23}$  particles.) Consider each situation below as it relates to the reaction in Model 1:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ .

a. Calculate the amount of reactants consumed and products made.

b. Record the ratio of  $\text{N}_2$  to  $\text{H}_2$  to  $\text{NH}_3$ . Reduce the ratio to the lowest whole number possible.

|                                                                                                         | $\text{N}_2$<br>Consumed | $\text{H}_2$<br>Consumed | $\text{NH}_3$<br>Produced | Ratio<br>$\text{N}_2:\text{H}_2:\text{NH}_3$ |
|---------------------------------------------------------------------------------------------------------|--------------------------|--------------------------|---------------------------|----------------------------------------------|
| If the reaction occurred $6.02 \times 10^{23}$ times, how many molecules would be consumed or produced? | $6.02 \times 10^{23}$    | $1.81 \times 10^{24}$    | $1.20 \times 10^{24}$     | 1 : 3 : 2                                    |
| How many <i>moles</i> of each substance would be consumed or produced in the previous situation?        | 1                        | 3                        | 2                         | 1 : 3 : 2                                    |

5. Refer to the data table in Question 4.

a. How do the reduced ratios in the last column compare to the coefficients in the reaction in Model 1?

*The reduced ratios are the same as the coefficients.*

b. Use mathematical concepts to explain how your answer in part a is possible.

*We are still multiplying coefficients by the same number so we just didn't everything by the number to get back to a reduced ratio.*

6. The ratio obtained from the coefficients in a balanced chemical equation is called the **mole ratio**.

a. What is the mole ratio for the reaction in Model 1?



b. Explain why this ratio is called the mole ratio?

*It is the reduced ratio in moles (all the coefficients =  $6.02 \times 10^{23}$ ).*

# Pogil Mole Ratio

**Liyong Dong**



**Pogil Mole Ratio:**

## **Pogil Mole Ratio** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Pogil Mole Ratio**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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