

Exercises

- Based on the data in Table 1, predict the boiling points of
 - heptane, $\text{CH}_3(\text{CH}_2)_5\text{CH}_3$
 - ethanol, $\text{CH}_3\text{CH}_2\text{OH}$
 - 2-octanone, $\text{CH}_3\overset{\text{O}}{\underset{\parallel}{\text{C}}}(\text{CH}_2)_5\text{CH}_3$
- Both *cis*-1,2,-dichloroethylene and *trans*-1,2,-dichloroethylene have the same molecular formula: $\text{C}_2\text{H}_2\text{Cl}_2$. However, the *cis* compound has a dipole moment, while the *trans* compound does not. One of these species has a boiling point of 60.3°C and the other has a boiling point of 47.5°C . Which compound has which boiling point? Explain your reasoning.
- Rank each of the following groups of substances in order of increasing boiling point, and explain your reasoning:
 - NH_3 , He, CH_3F , CH_4
 - CH_3Br , Ne, CH_3OH , CH_3CN
 - CH_4 , SiH_4 , GeH_4 , SnH_4
- Using grammatically correct English sentences, describe the difference between the hydrogen bond between two water molecules and the O–H bond in a particular water molecule.
- Fluoromethane, CH_3F , and methanol, CH_3OH , have approximately the same molecular weight. However, the boiling point of CH_3OH is 65.15°C , whereas the boiling point of CH_3F is almost 100 degrees lower, -78.4°C . Explain.

Problems

- You are given an unknown liquid to identify. You are told that the molecular formula of the compound is $\text{C}_2\text{H}_6\text{O}_2$. You measure the boiling point of the compound and find it to be 198°C . Identify this unknown liquid and explain your reasoning. You may wish to consider the following boiling points of various molecules in your analysis:

Pogil Chemactivity 27 Intermolecular Forces Activity

H Kauffman



Pogil Chemactivity 27 Intermolecular Forces Activity:

Pogil Chemactivity 27 Intermolecular Forces Activity Book Review: Unveiling the Magic of Language

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