INTERMOLECULAR FORCES - ANSWER KEY (from Tro, Chapter 11, page 512)

| | | ************* | -, | 7.0 | |
|-----|---|--|---|--|--|
| 49. | Determine the kinds of intermolecular forces that are present in each of the following elements or compounds: | | | | |
| | (a) Kr* | (b) NCI ₃ ** | (c) SiH ₄ * | (d) HF*** | |
| | (e) N ₂ * | (f) NH ₃ *** | (g) CO** | (h) CCl ₄ * | |
| 50. | Determine the kinds of intermolecular forces that are present in each of the following elements or compounds: | | | | |
| | (a) HCI** | (b) H ₂ O**** | (c) Br ₂ * | (d) He* | |
| | (e) PH ₃ ** | (f) HBr | (g) CH ₃ OH**** | (h) I ₂ " | |
| | * dispersion forces only; ** dispersion and dipole-dipole forces; *** dispersion force, dipole-dipole forces, and hydrogen bonds | | | | |
| 51. | Arrange the following in order of increasing boiling point. Explain your reasoning. | | | | |
| | (a) CH ₄ | (b) CH ₃ CH ₃ | (c) CH ₃ CH ₂ CI | (d) CH ₂ CH ₂ OH a b c c | |
| 52. | Arrange the following | in order of increasing boilin (b) H ₂ Se | ing point. Explain your reasoning. (c) H₂O a b c | | |
| 53. | For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning. (a) CH ₃ OH or CH ₃ SH (b) CH ₃ OCH ₃ or CH ₃ CH ₂ OH | | | | |
| | (c) CH ₄ or CH ₃ CH ₃ | | | | |
| 54. | For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning. | | | | |
| | (a) NH ₃ or CH ₄ (b) CS ₂ or CO ₂ (c) CO ₂ or NO ₂ | | | | |
| 55. | For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning. | | | | |
| | (a) Br ₂ or I ₂ | (b) H ₂ S or H ₂ O | (c) NH ₃ or PH ₃ | | |
| 56. | For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning. (a) CH ₂ or CH ₃ CI | | | | |
| | (b) CH ₃ CH ₂ CH ₂ OH or CH ₃ OH | | | | |
| | (c) CH₃OH or H₃CO | | | | |
| 57. | Which of the following pairs of substances would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved. | | | | |
| | (a) CCI₄ and H₂O | | (b) KCI and H ₂ O (ion-dipole forces) | | |
| | (c) Br ₂ and CCl ₄ (dispersion forces) (d) CH ₃ CH ₂ OH and | | | d H ₂ O (H-bonds) | |
| 58. | Which of the following pairs of compounds would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved. | | | | |
| | (a) CH ₂ CH ₂ CH ₂ CH ₃ and CH ₂ CH ₂ CH ₂ CH ₂ CH ₃ (dispersion forces) | | | | |
| | (b) CBr₄ and H₂O | | | | |
| | (c) LiNO ₃ and H ₂ O (ion-dipole forces) | | | | |
| | (d) CH ₃ OH and CH ₃ CH ₂ CH ₂ CH ₃ | | | | |

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