

Percent Yield Tutorial

PRACTICE PROBLEMS:

1. Industrial processes are especially concerned with percent yield. (Why?) For producing phosphoric acid the following reaction is used: $\text{P}_2\text{O}_5(\text{s}) + 3\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{H}_3\text{PO}_4(\text{aq})$. What is the expected mass of H_3PO_4 that will be produced from 50 grams of P_2O_5 reacting in excess water?

$$50 \text{ g } \text{P}_2\text{O}_5 \times \frac{1 \text{ mol}}{142 \text{ g } \text{P}_2\text{O}_5} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ P}_2\text{O}_5} \times \frac{98 \text{ g } \text{H}_3\text{PO}_4}{1 \text{ mol}} = 69.0 \text{ g } \text{H}_3\text{PO}_4$$

expected
⇒ theoretical yield

2. For the above problem, what if 65 grams of H_3PO_4 is actually produced. What is the % yield?

—CrashChemistry—

Percentage Yield Lab Grade 1

Judd E. Hollander

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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