- Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.
- a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)

- b. What type of reaction is this? single replacement
- c. If 16.2 g of Mg are heated with 12.0 g of H<sub>2</sub>O, which is the limiting reactant?

d. How many moles of the excess reactant remain after the reaction is complete? 16.29 Mgx \_ Love Mg x \_ 2mot H20 \_ 1.3 mol H20

e. How many grams of each product are formed?

e. You perform this reaction, and obtain 16.0 g Mg(OH)<sub>2</sub>. What is your percent yield?

# **Sample Problems H Gas Stoichiometry Answers**

**S Ashworth** 

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