

Vapour Pressure Curves

DEFINITION

The vapour pressure of a substance depends on the temperature (Boyle's temperature) due to the liquid-vapour pressure equilibrium. A liquid boils when the vapour pressure equals the atmospheric pressure. Water boils at 100°C at sea level but the pressure varies with altitude so at higher height and low baric atmospheric pressure it boils at low level. Water boils at a different temperature. Various substances as well as their solids and gases that equally boil at the same height vapour pressure.

Learning Objectives

1. Understand the relationship between vapour pressure and temperature.
2. Understand the vapour pressure of a substance in the boiling point.
3. Understand vapour pressure curve as function of height in strength of the temperature curve between the altitude of condensation.

Basic Concepts

1. Understand the pressure within description of vapour pressure and its relation.
2. Understand the boiling point of liquid from vapour pressure curve.
3. Understand various strength of the condensation curve based on data for vapour pressure curve.

Assessment

1. Understand the concept of vapour pressure, boiling point.

Prerequisites

1. Vapour pressure.
2. Pressure & temperature.
3. The boiling point.

Concepts and Relationship

Understand the relationship between the pressure and altitude.

1. Vapour pressure.
2. Condensation.
3. Boiling.
4. Chemical boiling point.
5. Under condensation degree of condensation.

Pogil Vapor Pressure Curves Answer Key

DP Hallahan



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