# Mole-Mole Stoichiometry Worksheet

Complete each of the following mole to mole stoichiometric conversions.

29) Given the following equation: 2 K + Cl<sub>2</sub> ---> 2 KCl How many moles of KCl are produced from 2.00 moles of K and excess Cl<sub>2</sub>?

30)Given the following equation: 8 Fe + S<sub>8</sub>---> 8 FeS How many moles of iron are needed to react with 16.0 moles of sulfur?

31)Given the following equation: Cu + 2 AgNO<sub>3</sub> ---> Cu(NO<sub>3</sub>)<sub>2</sub> + 2 Ag How many moles of Cu are needed to react with 3.00 moles of AgNO<sub>3</sub>?

32) Given the following equation: 2 K + Cl<sub>2</sub>---> 2 KCl How many moles of K are needed to react with 0.5 moles of Cl<sub>2</sub>?

33) 4 Fe + 3 O<sub>2</sub> → 2 Fe<sub>2</sub>O<sub>3</sub>
How many moles of Fe are needed to react with 1.00 moles of O<sub>2</sub>?

34) 2 Zn + O₂ → 2 ZnO
How many moles of zinc are needed to make 6 moles of zinc oxide?

35) 2 H₂ + O₂ → 2 H₂O
5.00 moles of hydrogen react with excess oxygen. What is the maximum number of moles of water that can be formed?

36). N₂ + 2 O₂ → N₂O₄
4 moles of oxygen react with excess nitrogen. What is the maximum number of moles of dinitrogen tetroxide (N₂O₄) that can be produced?

Given the following equation: 2 K + Cl<sub>2</sub> --- > 2 KCl
 How many grams of KCl are produced from 78.20 g of K and excess Cl<sub>2</sub>?

# **Pogil Mole Mole Stoichiometry**

Shasha Hu



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