

Information:

Percent Composition, also called % composition or % mass, is the percentage of each atom in a molecule by mass.

For example

$$\% \text{ Composition of Hydrogen} = \frac{\text{mass of all hydrogen in molecule/sample}}{\text{total mass of molecule/sample}}$$

You would repeat this for every atom in the molecule

3) Each group member pick 2 different hydrocarbons on the table, calculate the % composition of each element in each hydrocarbon. Compare answers within your group and fillout the remaining blanks in Table 1.

Ex) C_2H_2 mass = 26g/mol

$$C = 24g/mol \Rightarrow \% C = \frac{24g/mol}{26g/mol} = 92.3\%$$

$$H = 2g/mol \Rightarrow \% H = \frac{2g/mol}{26g/mol} = 7.7\%$$

4) Take a closer look at the hydrocarbons now, are there any that have a similar % composition? Are their properties similar as well?

C_2H_2 + C_6H_6 have the same % composition, but very different properties

C_3H_6 + C_4H_8 have the same % comp. but very different properties.

Just b/c molecules hv the same ratio does not mean they'll have same properties

Pogil Ap Gases

M Tight



Pogil Ap Gases:

A dictionary of the English language. To which are added, a synopsis of words differently pronounced and Walker's Key to the classical pronunciation of Greek, Latin and Scripture proper names. Revised and enlarged, by C.A. Goodrich Noah Webster, 1866 An American Dictionary of the English Language Noah Webster, 1832 **Gas Transport in Porous Media** E. A. MASON, 1983

The Properties of Gases Marylou Morano-Kjelle, 2006-08-15 Introduces the physical properties of gases and discusses why oxygen nitrogen and hydrogen are such important gases The Physical Properties of Gases Arthur Lalanne Kimball, 1890 **The Nature of a Gas** Leonard Benedict Loeb, 1931 Introduces the tractor trailer allowing the reader to count its parts using the numbers from one to ten and its wheels using the numbers to eighteen

The Laws of Gases: Memoirs Carl Barus, Robert Boyle, Émile Hilaire Amagat, 2023-07-18 This scientific memoir covers the fundamental laws and properties of gases as understood by three pioneering researchers in the field Through experiments and rigorous analysis the authors explore topics such as temperature pressure heat and chemical reactions shedding light on the complex and fascinating dynamics of gas molecules With detailed descriptions and illustrations this book provides a comprehensive introduction to the subject for students and scholars alike This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant **Gases**

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with applications of the soluble nature of gases **The Laws of Gases** Robert Boyle, Emile Hilaire Amagat, 1899

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straightforward diagrams and easy to understand text Readers will comprehend such physical science concepts as

evaporation condensation compressibility fluidity and volume with ease **Gases** Robin Nelson, 2005-01-01 Introduces

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