

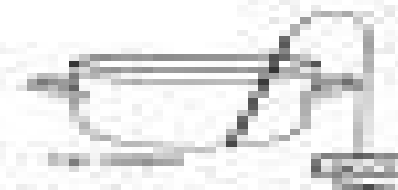
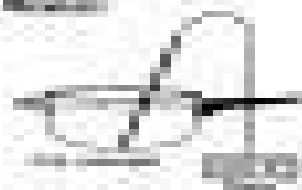
Calorimetry

What is calorimetry? How does it work? What is its purpose?

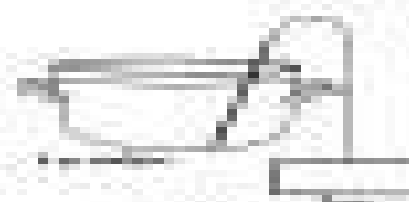
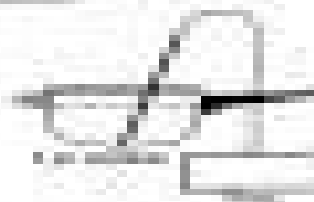
When a substance is heated, the temperature of that substance increases. Will the same amount of energy create different substances to have identical temperature increases? Will the same amount of energy be needed to create identical temperature increases in different amounts of the same substance? In this practice you will explore how mass, temperature, heat energy, and the type of substance are related.

Model 1 – A Plot of Mass

Calorimetry



How much energy is needed to heat 100g of water per second?



1. In Model 1, which substance had the most grams of water? 200g

2. Compare the graphs/curves in Model 1.

a. How much total joules of energy were added to the temperature? 100 J/g

b. How much total joules of energy were added to the temperature? 200 J

3. In Model 1, which substance had the highest mass? Which substance had the lowest mass? Which substance had the highest temperature increase? Which substance had the lowest temperature increase?

How much energy is needed to heat 100g of water per second? 100 J/g
 How much energy is needed to heat 200g of water per second? 200 J
 How much energy is needed to heat 100g of water per second? 100 J/g
 How much energy is needed to heat 200g of water per second? 200 J

Pogil Activities For High School Chemistry Calorimetry Answers

Sabine Zange



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