Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)



Atomic Rodii Size:

1. Explain atomic radius: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.

2. Using the information in figure 5-13 (pg 141) in your book, write the atomic radii size for alkali metals & period #2 on the periodic table to your right.



- 3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
- 4. Of all the elements on the periodic table, identify & circle the largest & smallest in size.
- Rank the following elements from largest to smallest in atomic radii: a) Be, Ba, Mg = Ba, Mg, Be
- 6. How can you explain the trend:

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on

in creases the attraction of the nucleus and therefore cloud close to the nucleus resulting

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R Sandford

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