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#### Exceed to Succeed

## **Worksheet: Periodic Trends**

#### 1. ATOMIC RADIUS

For each of the following sets of atoms, rank the atoms from smallest to largest atomic radius.

- a. Li, C, F
- b. Li. Na. K
- c. Ge. P. O.
- d. C. N. Al
- e. Al. Cl. Ga.

#### IONIC RADIUS 2...

For each of the following sets of ions, rank them from smallest to largest ionic radius.

- a. Mg<sup>2+</sup>, Si<sup>4-</sup>, S<sup>2-</sup> b. Mg<sup>2+</sup>, Ca<sup>2+</sup>, Ba<sup>2+</sup>
- c. F. Cl. Br
- d. Ba<sup>2+</sup>, Cu<sup>2+</sup>, Zn<sup>2+</sup> e. Si<sup>4-</sup>, P<sup>3-</sup>, O<sup>2-</sup>

#### IONIZATION ENERGY 3...

For each of the following sets of atoms, rank them from lowest to highest ionization energy.

- a. Mg, Si, S
- b. Mg, Ca, Ba
- c. F. Cl. Br
- d. Ba. Cu. Ne
- e. Si. P. He

#### 46. ELECTRONEGATIVITY

For each of the following sets of atoms, rank them from lowest to highest electronegativity.

- a. Li, C, N
- b. C. O. Ne
- c. Si. P. O
- d. K. Mg, P
- e. S. F. He.

# <u>Periodic Trends And Properties Of Elements Lab Answer</u> <u>Keys</u>

**Pam Walker, Elaine Wood** 

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