

Chemistry I: Periodicity Worksheet

- 1) Arrange the following elements in order of increasing size: Al, B, C, K, and Na

Background

- 2) Arrange the following elements in order of increasing size: Ca, Rb, P, Ge, and Sr.

Abstract

33. Circle the atom or ion in each pair that has the larger radius.

- ch In or I

- 4) Select the atom or ion in each pair that has the largest radius.

- ch. Error: As

- 5) Which of the following groups of elements is arranged correctly in order of increasing ionization energy?
(Answer by circling the correct answer.)

- (d)
- $\text{Ne} < \text{C} < \text{Si} < \text{Li}$

- 6) Arrange the following atoms in order of increasing ionization energy: Li, K, C, and N.

1000

- 7) Arrange the following atoms in order of increasing ionization energy: Si, K, As, and Ca.

1000

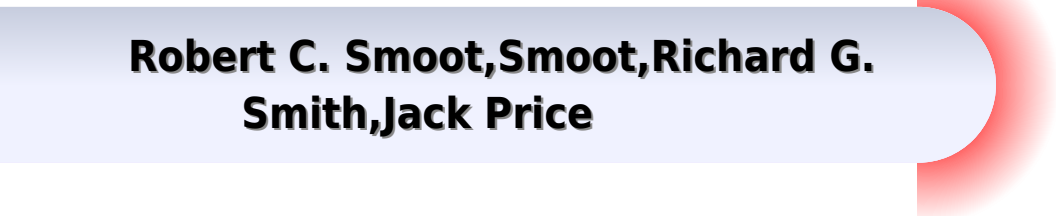
- 8) Compare the elements Na, Mg, O, and P.

- c) Place the elements in order of increasing ionization energy.

1999

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**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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