

Step 8. Periodic table trends

Electronegativity

Ionization energy

He

Atomic radius

The periodic table displays various elements with their symbols and atomic numbers. Overlaid on the table are several arrows indicating trends:

- Electronegativity:** A black arrow pointing from left to right across the top of the table.
- Ionization energy:** A green arrow pointing from left to right across the top of the table, and a green arrow pointing from bottom to top on the right side of the table.
- Atomic radius:** A red arrow pointing from top to bottom on the left side of the table, and a red arrow pointing from right to left across the bottom of the table.
- Non-metallic properties:** A blue arrow pointing from bottom-left to top-right, passing through the upper right quadrant of the table.
- Metallic properties:** A light blue arrow pointing from top-left to bottom-right, passing through the lower left quadrant of the table.

Ionization Energy

Electronegativity

Atomic radius

An arrow head shows the INCREASE of the property

Example: comparing K atom with Br atom

- Atomic radius of K is HIGHER than atomic radius of Br
- Ionization energy of K is LOWER than ionisation energy of Br
- Metallic properties of K are HIGHER than metallic properties of Br



Periodic Trends Answer Key Section Review

Clifford Lane



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