

NAME \_\_\_\_\_

DATE \_\_\_\_\_

### GRAPHING PERIODIC TRENDS

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.

Define the following terms(starting on page 170):

Atomic radius-

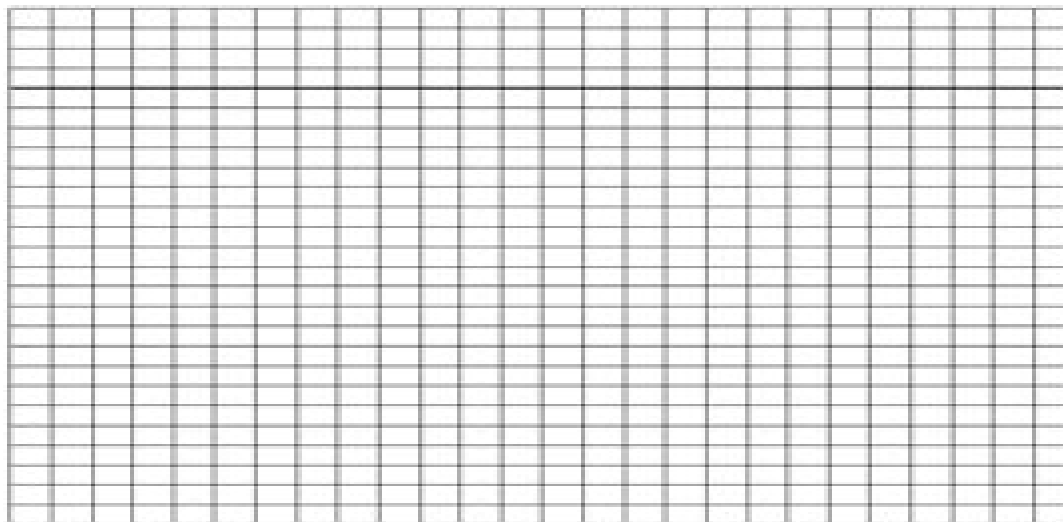
First ionization energy-

Electronegativity-

#### Graph 1 – Atomic Radius as a function of Atomic Number

- A. Create a graph of the atomic radius as a function of atomic number. Plot atomic number on the X axis and atomic radius on the Y axis. Remember to label the axes!
- B. Use a colored pen, pencil or highlighter to **trace** over the element's period (horizontal row on the periodic table). For example: use GREEN to **trace** for all of the elements in row 1, then use YELLOW to **trace** for all of the elements in row 2, then use ORANGE to **trace** for all the elements in row 3, then use BLUE to **trace** for all the elements in row 4.

| Symbol | Atomic Number | Atomic Radius(pm) |
|--------|---------------|-------------------|
| H      | 1             | 37                |
| He     | 2             | 50                |
| Li     | 3             | 123               |
| Be     | 4             | 89                |
| B      | 5             | 80                |
| C      | 6             | 77                |
| N      | 7             | 70                |
| O      | 8             | 66                |
| F      | 9             | 64                |
| Ne     | 10            | 67                |
| Na     | 11            | 157               |
| Mg     | 12            | 136               |
| Al     | 13            | 125               |
| Si     | 14            | 117               |
| P      | 15            | 110               |
| S      | 16            | 104               |
| Cl     | 17            | 99                |
| Ar     | 18            | 98                |
| K      | 19            | 203               |
| Ca     | 20            | 174               |



1. Describe the trend in atomic radius as the atomic number increases across a period.
2. Describe the trend in atomic radius as the atomic number increases down a group.

# Periodic Trends Answers Extension Questions

**Tina Overton, Jonathan Rourke, Fraser  
A. Armstrong**



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Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998

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New York Math: Math B ,2000

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