

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Calculate:** Moles of product based on the limiting reagent. **Remember:** Percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$

Chemists will determine the limiting reagent by calculating the moles of product that each reactant will form. The reactant that forms the least moles of product is the limiting reagent.



The reactant required with the smallest amount is the limiting reagent.

2H ₂ (g)	100.00 grams
O ₂ (g)	50.00 grams
H ₂ O(l)	100.00 grams
H ₂ (g)	100.00 grams

- 100.00 g of each reactant will be used to create 100.00 g of product from 100.00 g of reactants. **Remember:** Limiting reagent will be the reactant that will create the least product.



- 100.00 g of H₂(g) will create 100.00 g of H₂O(l) and 100.00 g of O₂(g) will create 112.50 g of H₂O(l).

- 100.00 g of H₂(g) will create the least amount of product, so it is the limiting reagent.



- 100.00 g of H₂(g) will create 100.00 g of H₂O(l) and 100.00 g of O₂(g) will create 112.50 g of H₂O(l). **Remember:** 100.00 g of H₂(g) will be the limiting reagent and 100.00 g of H₂O(l) will be the actual yield.



Percent Yield And Limiting Reactants Quiz

Emilie Sanchez



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