

Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.

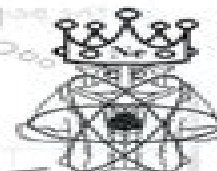


Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Fluorine
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (column 17), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



I am so perfect... I have a full valence shell. I don't need those other elements... I hope they just leave me alone!



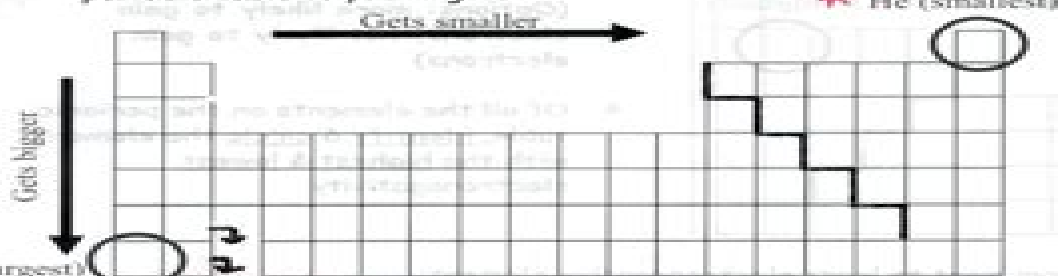
Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

Effective nuclear charge, Z_{eff} , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.

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4AP Answer

Periodic Trends Answers Key Pogil

LL Leslie



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Chemical Periodicity Robert Thomas Sanderson, 2013-04-20 **Chemistry and Our Universe**, Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there

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The Amazing World of the Atom and the Ultimate Periodic Table of the Elements Costel Rizescu, Mihaela Rizescu, 2018-05-24 This book is addressed to undergraduate and graduate students taking physics and chemistry courses at an introductory level STEM high school students and to more experienced professors and researchers who need to stay up to date with the most recent advances in data as well It is also beneficial to those involved in materials science medical physics biotechnology astronomy quantum chemistry Earth and space science and other related fields The volume consists of two parts the first part is devoted to the quantum mechanics theory of atomic structures as discovered and verified by physics whereas the second part presents and analyzes a large amount of the latest critically evaluated data for the properties of all elements on the periodic table of the elements Periodic trends for ionization energy electronegativity atomic radius density and melting points are graphed and thoughtfully analyzed Additional data of ionic radius covalent radius van der Waals radius and 12 coordination atomic radius for almost all

elements on the periodic table are also presented and evaluated The authors discussed over 30 problems review questions and their solutions to help deepen the insight in this subject area to make this book a real study text Plenty of problems are given to elucidate the material This book represents the culmination of the authors many years of research and teaching physics materials science and instrumentation It is also a crystallization of their passion and strong interest in the history of atomic physics and the philosophy of science in general From the beginning it was authors intention to include information in this book that is not easily located elsewhere or is not found completely in a single book or publication Thus while this book can be used as a text the authors hope that it will be used as a useful reference too Each chapter and sub chapter of this book capitalizes on the strengths comments feedback and criticism that the authors expect to have from students faculty and working professionals

The Periodic Table Eric R. Scerri, 2019-07-25 The periodic table of elements first encountered by many of us at school provides an arrangement of the chemical elements ordered by their atomic number electron configuration and recurring chemical properties and divided into periodic trends In this Very Short Introduction Eric R Scerri looks at the trends in properties of elements that led to the construction of the table and shows how the deeper meaning of the table s structure gradually became apparent with the development of atomic theory and in particular quantum mechanics which underlies the behaviour of all of the elements and their compounds This new edition publishing in the International Year of the Periodic Table celebrates the completion of the seventh period of the table with the ratification and naming of elements 113 115 117 and 118 as nihonium moscovium tennessine and oganesson Eric R Scerri also incorporates new material on recent advances in our understanding of the origin of the elements as well as developments concerning group three of the periodic table

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