

The Periodic Table

The Structure of the Periodic Table

- The periodic table displays the elements in order of increasing atomic number i.e. in order of increasing number of protons.
- Elements are arranged in 18 vertical columns called groups that are numbered 1 – 18.
- The seven horizontal rows in periodic table are called periods. The periods are numbered 1-7. The first period contains only two elements, hydrogen and helium.
- Elements within the same groups have similar chemical properties because they have the same number of valence electrons. The number of valence electrons is the same as the group number (in Roman numerals).

PERIODS 1-7		GROUPS 1-18																				
																		17	18			
		1																	17	18		
		2	Li	Be													B	C	N	O	F	Ne
		3	Na	Mg													Al	Si	P	S	Cl	Ar
		4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr		
		5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe		
		6	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn		
		7	Fr	Ra	Ac	Rf	Ha	Sg	Nh	Hs	Mt	Dns	Dns									

Energy levels and period numbers

The electrons of an atom are arranged in energy levels. The period number tells us in which energy level an atom has its valence electrons.

For example:

Element	In which Group?	In which Period?	How many valence electrons	Which energy level are valence electrons?
Phosphorous	15	3	5	3 rd energy level
Potassium				
Calcium				
Oxygen				

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Clemens Wendtner



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