

Percent Yield Tutorial

PRACTICE PROBLEMS:

1. Industrial processes are especially concerned with percent yield. (Why?) For producing phosphoric acid the following reaction is used: $\text{P}_2\text{O}_5(\text{s}) + 3\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{H}_3\text{PO}_4(\text{aq})$. What is the expected mass of H_3PO_4 that will be produced from 50 grams of P_2O_5 reacting in excess water?

$$50 \text{ g } \text{P}_2\text{O}_5 \times \frac{1 \text{ mol}}{142 \text{ g } \text{P}_2\text{O}_5} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ P}_2\text{O}_5} \times \frac{98 \text{ g } \text{H}_3\text{PO}_4}{1 \text{ mol}} = 69.0 \text{ g } \text{H}_3\text{PO}_4$$

expected
⇒ theoretical yield

2. For the above problem, what if 65 grams of H_3PO_4 is actually produced. What is the % yield?

—CrashChemistry—

Percentage Yield Lab Grade 1

Michael Hoelscher

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Michael Hoelscher.

Percentage Yield Lab Grade 1:

Percentage Yield Lab Grade 1 Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Percentage Yield Lab Grade 1**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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